

Unit 6 – Worksheet 3

Ionic Compounds

Properties

Ionic compounds generally have higher m.p's and b.p's than do molecular compounds. When molten, they conduct electricity. If soluble, the solutions also conduct electricity.

Basic structural unit

The basic structural unit is the formula unit, the simplest ratio of atoms in a crystalline ionic compound.

1. Give the name of the following simple binary ionic compounds.

- a. Na_2O sodium oxide
- b. K_2S potassium sulfide
- c. MgCl_2 magnesium chloride
- d. CaBr_2 calcium bromide
- e. BaI_2 barium iodide
- f. Al_2S_3 aluminum sulfide
- g. CaBr cesium bromide
- h. AgF silver fluoride

2. Give the name of the following simple binary ionic compounds.

- a. Na_3N sodium nitride
- b. K_2O potassium oxide
- c. AgBr silver bromide
- d. MgI_2 magnesium iodide
- e. SrO strontium oxide

3. Write the formula for the following binary ionic compounds.

- a. lithium bromide LiBr
- b. sodium iodide NaI
- c. silver sulfide Ag_2S
- d. cesium oxide Cs_2O
- e. beryllium iodide BeI_2
- f. barium hydride BaH_2
- g. aluminum fluoride AlF_3
- h. potassium oxide K_2O

4. Write the formula for these ionic substances.

- a. silver oxide Ag_2O
- b. aluminum sulfide Al_2S_3
- c. sodium nitride Na_3N
- d. barium chloride BaCl_2
- e. strontium hydride SrH_2

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