



c) [NaOH] (in mol/L),

$$x \text{ mol NaOH} = 0.02500 \text{ L} \times \frac{0.00690 \text{ mol}}{\text{L}} \times \frac{1 \text{ mol NaOH}}{1 \text{ mol HA}} = 1.73 \times 10^{-4} \text{ mol NaOH}$$

$$[NaOH] = \frac{1.73 \times 10^{-4} \text{ mol}}{0.02135 \text{ L}} = \underline{\underline{0.0081 \text{ M}}}$$

d) if the resulting solution is acidic, basic or neutral at the equivalence point, and

A^- is weakly basic $\Rightarrow$ basic solution

e) an appropriate indicator for the titration.

Phenolphthalein

8. Triethanolamine, a weak monoprotic base, is titrated with HCl. The following data are obtained.

- 20.00 mL of triethanolamine
- 16.30 mL of HCl
- Initial pH of acid = 10.355

– $pH_{\frac{1}{2}} = 7.763$ at addition of 8.15 mL of acid during titration

$$pOH_{\frac{1}{2}} = 14 - 7.763 = 6.237$$

Determine

a) the K_b for triethanolamine,

$$K_b = [OH^-]^2$$

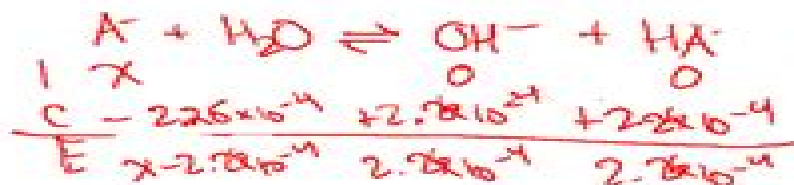
$$= 10^{-pOH_{\frac{1}{2}}}$$

$$= \underline{\underline{5.79 \times 10^{-7}}}$$

THIS should read initial [triethanolamine] $\star$

b) initial ~~10.355~~ (in mol/L).

$$pOH_{\text{init}} = 14.000 - 10.355 = 3.645$$



$$K_a = \frac{[OH^-][HA]}{[CH_3COOH]} = 5.79 \times 10^{-7}$$

$$x - 2.2 \times 10^{-4} = \frac{(2.26 \times 10^{-4})^2}{5.79 \times 10^{-7}}$$

$$x = 0.0859 + 0.000212$$

$$\underline{\underline{x = 0.0861 \text{ M}}}$$

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David Kirk



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Table of Contents Modeling Chemistry U6 Review Vanswers

1. Understanding the eBook Modeling Chemistry U6 Review Vanswers
 - The Rise of Digital Reading Modeling Chemistry U6 Review Vanswers
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry U6 Review Vanswers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry U6 Review Vanswers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry U6 Review Vanswers
 - Personalized Recommendations
 - Modeling Chemistry U6 Review Vanswers User Reviews and Ratings
 - Modeling Chemistry U6 Review Vanswers and Bestseller Lists

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 - Modeling Chemistry U6 Review Vanswers Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
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 - Modeling Chemistry U6 Review Vanswers Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Modeling Chemistry U6 Review Vanswers
 - Highlighting and Note-Taking Modeling Chemistry U6 Review Vanswers
 - Interactive Elements Modeling Chemistry U6 Review Vanswers
8. Staying Engaged with Modeling Chemistry U6 Review Vanswers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry U6 Review Vanswers
9. Balancing eBooks and Physical Books Modeling Chemistry U6 Review Vanswers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry U6 Review Vanswers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modeling Chemistry U6 Review Vanswers
 - Setting Reading Goals Modeling Chemistry U6 Review Vanswers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modeling Chemistry U6 Review Vanswers
 - Fact-Checking eBook Content of Modeling Chemistry U6 Review Vanswers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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