

Model 1: Rules of Solubility in Aqueous Solutions

Several solid compounds are placed in water to determine if they are soluble. The results are shown in the chart below. If it does dissolve no mark is made.

The top row shows the cation of the compound. The far left column shows the anion of the compound. For example, $\text{Mg}(\text{OH})_2$ is insoluble, MgBr_2 is soluble and AgCl is insoluble.

Table 1 – This table presents an overview of solubility of selected salts in water

X = does not dissolve in water. * partially soluble.

	NH_4^+	Li^+	Na^+	K^+	Mg^{2+}	Ca^{2+}	Sr^{2+}	Ba^{2+}	Ag^+	Pb^{2+}	Hg_2^{2+}	Fe^{3+}	Cu^{2+}	Zn^{2+}
NO_3^-														
$\text{C}_2\text{H}_3\text{O}_2^-$														
CO_3^{2-}														
$\text{C}_2\text{O}_4^{2-}$														
Cl^-									X	X	X			
Br^-									X	X	X			
I^-									X	X	X			
SO_4^{2-}						X	X	X	X	X	X			
OH^-					X	*	*		X	X	X	X	X	X
S^{2-}					X	X	X		X	X	X	X	X	X
CO_3^{2-}					X	X	X	X	X	X	X	X	X	X
PO_4^{3-}					X	X	X	X	X	X	X	X	X	X
$\text{C}_2\text{O}_4^{2-}$					X	X	X	X	X	X	X	X	X	X
CrO_4^{2-}					X	X	X	X	X	X	X	X	X	X

Key Questions

1. Is calcium carbonate soluble or insoluble?
2. Is silver bromide soluble or insoluble?
3. Is iron (III) sulfate soluble or insoluble?
4. For what cations are the compounds always soluble in water?
5. For what anions are the compounds always soluble in water?
6. For what anions are most of the compounds usually soluble (with some exceptions)?
7. For those anions that usually form soluble compounds, which cations result in the

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Table of Contents Model Rules Of Solubility In Aqueous Solutions Answers

1. Understanding the eBook Model Rules Of Solubility In Aqueous Solutions Answers
 - The Rise of Digital Reading Model Rules Of Solubility In Aqueous Solutions Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Model Rules Of Solubility In Aqueous Solutions Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Model Rules Of Solubility In Aqueous Solutions Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Model Rules Of Solubility In Aqueous Solutions Answers

- Personalized Recommendations
 - Model Rules Of Solubility In Aqueous Solutions Answers User Reviews and Ratings
 - Model Rules Of Solubility In Aqueous Solutions Answers and Bestseller Lists
5. Accessing Model Rules Of Solubility In Aqueous Solutions Answers Free and Paid eBooks
 - Model Rules Of Solubility In Aqueous Solutions Answers Public Domain eBooks
 - Model Rules Of Solubility In Aqueous Solutions Answers eBook Subscription Services
 - Model Rules Of Solubility In Aqueous Solutions Answers Budget-Friendly Options
 6. Navigating Model Rules Of Solubility In Aqueous Solutions Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Model Rules Of Solubility In Aqueous Solutions Answers Compatibility with Devices
 - Model Rules Of Solubility In Aqueous Solutions Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Model Rules Of Solubility In Aqueous Solutions Answers
 - Highlighting and Note-Taking Model Rules Of Solubility In Aqueous Solutions Answers
 - Interactive Elements Model Rules Of Solubility In Aqueous Solutions Answers
 8. Staying Engaged with Model Rules Of Solubility In Aqueous Solutions Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Model Rules Of Solubility In Aqueous Solutions Answers
 9. Balancing eBooks and Physical Books Model Rules Of Solubility In Aqueous Solutions Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Model Rules Of Solubility In Aqueous Solutions Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Model Rules Of Solubility In Aqueous Solutions Answers
 - Setting Reading Goals Model Rules Of Solubility In Aqueous Solutions Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Model Rules Of Solubility In Aqueous Solutions Answers

- Fact-Checking eBook Content of Model Rules Of Solubility In Aqueous Solutions Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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