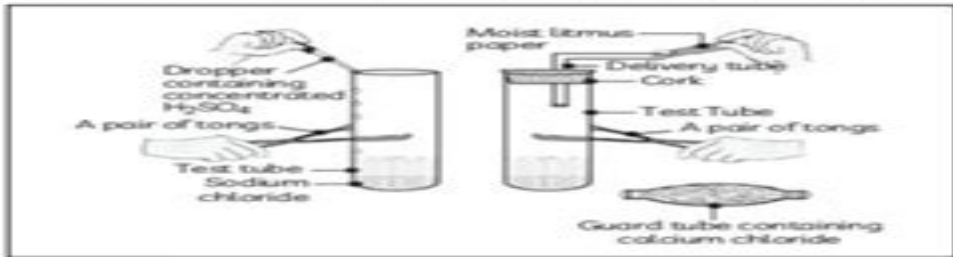
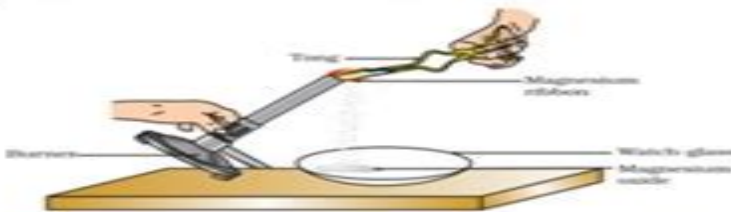


SECTION - A

Select and write one most appropriate option out of the four options given for each of the questions 1 – 20

Q. No	Questions	Marks																									
1	The change in colour of the moist litmus paper in the given set up is due to  i. presence of acid ii. presence of base iii. presence of H⁺(aq) in the solution iv. presence of Litmus which acts as an indicator (a) i and ii (b) Only ii (c) Only iii (d) Only iv.	1																									
2	In the redox reaction $MnO_2 + 4HCl \rightarrow MnCl_2 + 2H_2O + Cl_2$ (a) MnO_2 is reduced to $MnCl_2$ & HCl is oxidized to H_2O (b) MnO_2 is reduced to $MnCl_2$ & HCl is oxidized to Cl_2 (c) MnO_2 is oxidized to $MnCl_2$ & HCl is reduced to Cl_2 (d) MnO_2 is oxidized to $MnCl_2$ & HCl is reduced to H_2O	1																									
3	 Which of the following is the correct observation of the reaction shown in the above set up? (a) Brown powder of Magnesium oxide is formed. (b) Colourless gas which turns lime water milky is evolved. (c) Magnesium ribbon burns with brilliant white light. (d) Reddish brown gas with a smell of burning Sulphur has evolved.	1																									
4	With the reference to four gases CO_2, CO, Cl_2 and O_2, which one of the options in the table is correct? <table><tr><th>Option</th><th>Acidic oxide</th><th>Used in treatment of water</th><th>Product of respiration</th><th>Product of incomplete combustion</th></tr><tr><td>(a)</td><td>CO</td><td>Cl_2</td><td>O_2</td><td>CO</td></tr><tr><td>(b)</td><td>CO_2</td><td>Cl_2</td><td>CO_2</td><td>CO</td></tr><tr><td>(c)</td><td>CO_2</td><td>O_2</td><td>O_2</td><td>CO_2</td></tr><tr><td>(d)</td><td>CO</td><td>O_2</td><td>CO_2</td><td>CO_2</td></tr></table>	Option	Acidic oxide	Used in treatment of water	Product of respiration	Product of incomplete combustion	(a)	CO	Cl_2	O_2	CO	(b)	CO_2	Cl_2	CO_2	CO	(c)	CO_2	O_2	O_2	CO_2	(d)	CO	O_2	CO_2	CO_2	1
Option	Acidic oxide	Used in treatment of water	Product of respiration	Product of incomplete combustion																							
(a)	CO	Cl_2	O_2	CO																							
(b)	CO_2	Cl_2	CO_2	CO																							
(c)	CO_2	O_2	O_2	CO_2																							
(d)	CO	O_2	CO_2	CO_2																							

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Table of Contents Mock Test Papers For Class 10 Cbse

1. Understanding the eBook Mock Test Papers For Class 10 Cbse
 - The Rise of Digital Reading Mock Test Papers For Class 10 Cbse
 - Advantages of eBooks Over Traditional Books
2. Identifying Mock Test Papers For Class 10 Cbse
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mock Test Papers For Class 10 Cbse
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mock Test Papers For Class 10 Cbse
 - Personalized Recommendations
 - Mock Test Papers For Class 10 Cbse User Reviews and Ratings
 - Mock Test Papers For Class 10 Cbse and Bestseller Lists

5. Accessing Mock Test Papers For Class 10 Cbse Free and Paid eBooks
 - Mock Test Papers For Class 10 Cbse Public Domain eBooks
 - Mock Test Papers For Class 10 Cbse eBook Subscription Services
 - Mock Test Papers For Class 10 Cbse Budget-Friendly Options
6. Navigating Mock Test Papers For Class 10 Cbse eBook Formats
 - ePub, PDF, MOBI, and More
 - Mock Test Papers For Class 10 Cbse Compatibility with Devices
 - Mock Test Papers For Class 10 Cbse Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mock Test Papers For Class 10 Cbse
 - Highlighting and Note-Taking Mock Test Papers For Class 10 Cbse
 - Interactive Elements Mock Test Papers For Class 10 Cbse
8. Staying Engaged with Mock Test Papers For Class 10 Cbse
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mock Test Papers For Class 10 Cbse
9. Balancing eBooks and Physical Books Mock Test Papers For Class 10 Cbse
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mock Test Papers For Class 10 Cbse
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mock Test Papers For Class 10 Cbse
 - Setting Reading Goals Mock Test Papers For Class 10 Cbse
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mock Test Papers For Class 10 Cbse
 - Fact-Checking eBook Content of Mock Test Papers For Class 10 Cbse
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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