

Gas Law Worksheets Answer Key

Ws #1

- 106 mL
- 9.52 atm
- 1.00×10^2 torr
- 8.83 L
- 34.5 mL
- 9.90×10^2 torr
- 333 mL

Ws #2

- $V_2 = 3.47$ L
- $P_2 = 1970$ torr
- $P_2 = .910$ atm
- $V_2 = 180$ L
- $T_2 = 330$ K or 57°C
- $V_2 = 8.36$ L
- $P_2 = 4.36$ atm
- $V_2 = 4.9$ L
- $V_2 = 6250$ L
- $V_2 = 2.600$ L

Ws #3

- $V_2 = 104$ mL
- $P_2 = 394$ KPa
- $V_2 = 3.72$ L
- $P_2 = 1.0 \times 10^2$ torr
- $V_2 = 8.83$ L
- $V_2 = 34.5$ mL
- 700 mmHg (1 sf)
- $V_2 = .566$ L
- $V_2 = 87.2$ L
- $V_2 = 26.3$ mL

Ws#4

- 1.22 atm
- 6.17 L
- 101 g Cl_2
- 619 KPa
- .0157 mol O_2
- 81.5 g/mol
- 99.3 g/mol
- 13.9 g CO_2
- 33.7 atm
- 204 L

Ws#5

- 21.2 L H_2
- .272 g Mg
- $4\text{NH}_3 + 5\text{O}_2 \rightarrow 4\text{NO} + 6\text{H}_2\text{O}$
a) 18.8 L O_2 b) 5.3 mol H_2O c) 13.9 g H_2O
- 7.98 g LiOH

Ws#6

- $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$ b) 2.00 mol H_2O c) .500 mol H_2O
- a) 2-1-2-2 b) 2.5 L H_2
- a) 1-2-1-1 b) 11 L C_2H_2
- $2\text{Na}_2\text{O}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{NaOH} + \text{O}_2$ b) 6.40 g Na_2O_2
- a) 1-1-2-1 b) 186 L H_2
- $\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow \text{H}_2 + \text{ZnSO}_4$ b) 1.37g Zn

Ws#7

- NaN_3 , KNO_3 , SiO_2
- N_2
- 2-2-3, 10-2-1-5-1
- The first reaction produces more N_2 . You can tell by the coefficient.
- Na metal is very reactive, explosive
- 3.05 mol N_2 , 2.03 mol Na
- .203 mol N_2
- 3.255 mol N_2
- 72.87 L N_2
- 79.03 L
- 1.0×10^2 L
- So excess gas can escape.
- You would get way too much N_2 produce and it would hurt the driver.
- Yes ! ☺

Ws#8

- a) 2.00×10^2 mL, b) 630 mL, c) 2.5×10^2 mL, d) 1.00×10^3
- a) 1.17 atm, b) .916 atm, c) 1.1 atm d) 2.00 atm
- a) 29.2 mL b) 22.9 mL, c) 27 mL, d) 50 mL
- 2.12 atm
3. 2.06×10^{23} molecules
- 2.12 atm
- .0264 mol O_2
- 63.7°C
- 69.3 g/mol
- 85.2 mL
- 1.7 atm
- 2.01 L

Mixed Gas Laws Answer Key

Daniel F McAuley



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