

Stoichiometry Reflect #2

Learning Target... I can use stoichiometry to calculate amounts

1a. Balance this equation



1b. Determine how many moles of ammonium oxide was used if the student weighed out 55.2 g of it.

$$55.2 \text{ g (NH}_4\text{)}_2\text{O} \left(\frac{1 \text{ mol A}}{52.10 \text{ g}} \right) = \boxed{1.06 \text{ moles (NH}_4\text{)}_2\text{O}}$$

1c. Determine the amount of moles of ammonium nitride produced in this reaction.

$$1.06 \text{ moles A} \left(\frac{2 \text{ mol B}}{3 \text{ mol A}} \right) = \boxed{0.706 \text{ moles (NH}_4\text{)}_3\text{N}}$$

1d. Determine how many grams of ammonium nitride produced in this reaction.

$$0.706 \text{ moles} \left(\frac{108.11 \text{ g}}{1 \text{ mol B}} \right) = \boxed{48.1 \text{ g (NH}_4\text{)}_3\text{N}}$$

2. Using the following chemical reaction:

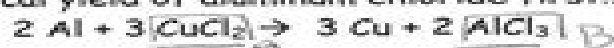


and, knowing that 61.6 g of copper was produced in this experiment, how many molecules of mercury nitrate was also produced?

$$61.6 \text{ g Cu} \left(\frac{1 \text{ mol A}}{63.55 \text{ g}} \right) \left(\frac{2 \text{ mol B}}{1 \text{ mol A}} \right) \left(\frac{6.02 \times 10^{23} \text{ mc}}{1 \text{ mol B}} \right) = \boxed{1.17 \times 10^{24} \text{ mc HgNO}_3}$$

3. Challenge...

What is the percent yield of the following reaction if 12.8 g of copper (II) chloride was used in the lab, and 7.75 g of aluminum chloride was measured at the end of the experiment? (hint: you must calculate the theoretical yield of aluminum chloride first!!)



$$12.8 \text{ g A} \left(\frac{1 \text{ mol A}}{134.45 \text{ g}} \right) \left(\frac{2 \text{ mol B}}{3 \text{ mol A}} \right) \left(\frac{133.33 \text{ g}}{1 \text{ mol B}} \right) = 8.46 \text{ g AlCl}_3$$

theoretical

$$\% \text{ yield} = \frac{7.75 \text{ g}}{8.46 \text{ g}} = \boxed{91.4\%}$$

Show work on back

Mixed Stoichiometry Practice Answer Key

Puja Mehta



Mixed Stoichiometry Practice Answer Key:

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In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Mixed Stoichiometry Practice Answer Key**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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