

24.9 mol

2. What volume would the CO_2 in Problem 1 occupy at STP?

$$24.9 \text{ mol} \times 22.4 = 557.76 \text{ L}$$

560 L

3. A sample of NH_3 gas occupies 75.0 liters at STP. How many molecules is this?

$$\frac{75.0 \text{ L}}{22.4} = 3.3482 \text{ mol} \times (6.02 \times 10^{23}) = 20.156 \times 10^{23}$$

4. What is the mass of the sample of NH_3 in Problem 3?

5. How many atoms are there in 1.3×10^{22} molecules of NO_2 ?

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R Bogdan



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