

Identification of a Metal Carbonate

Experiment 3: A study in stoichiometry

Metal carbonates react with hydrochloric acid to produce water and carbon dioxide gas. This is the "acid test" used by geologists to identify carbonate containing rocks and minerals. The equation for this reaction, as it pertains to this experiment, is:



or for a bicarbonate (hydrogen carbonate):



Carbon dioxide is fairly soluble in water under high pressure conditions, as evidenced in carbonated soft drinks, however at normal atmospheric pressure the carbon dioxide escapes and the carbonated beverage goes flat if left without being tightly capped.

In this experiment, a weighted sample of unknown metal carbonate will chemically react and the mass of carbon dioxide escaping will be determined. This will provide the data needed to calculate the molar mass of the unknown and determine which of the substances listed in Table 1 can be identified as the unknown metal carbonate. **WARNING!** Some of the unknowns as well as the hydrochloric acid used in this experiment are very corrosive substances. Wear Eye Protection and avoid skin contact. Wash with running water if exposed.

PRE-LAB ASSIGNMENT: In the space provided in Table 1, write the chemical name and molar mass for each of the following carbonates.

Table 1: Metal Carbonate Unknowns

CHEMICAL NAME	CHEMICAL FORMULA	MOLAR MASS
	Na₂CO₃	
	Na₂CO₃ · ½ H₂O	
	NaHCO₃	
	K₂CO₃	
	KHCO₃	
	K₂CO₃ · ½ H₂O	

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The Chemical News and Journal of Industrial Science William Crookes, James H. Gardiner, Gerald Druce, H. W. Blood-Ryan, 1897 **The Druggists' Circular and Chemical Gazette**, 1889 Includes Red book price list section title varies slightly issued semiannually 1897 1906 The Engineer, 1919 *Chemical News and Journal of Industrial Science*, 1897

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Table of Contents Metal Carbonate With Hydrochloric Acid Experiment Isa

1. Understanding the eBook Metal Carbonate With Hydrochloric Acid Experiment Isa
 - The Rise of Digital Reading Metal Carbonate With Hydrochloric Acid Experiment Isa
 - Advantages of eBooks Over Traditional Books
2. Identifying Metal Carbonate With Hydrochloric Acid Experiment Isa
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Metal Carbonate With Hydrochloric Acid Experiment Isa
 - User-Friendly Interface
4. Exploring eBook Recommendations from Metal Carbonate With Hydrochloric Acid Experiment Isa
 - Personalized Recommendations
 - Metal Carbonate With Hydrochloric Acid Experiment Isa User Reviews and Ratings
 - Metal Carbonate With Hydrochloric Acid Experiment Isa and Bestseller Lists
5. Accessing Metal Carbonate With Hydrochloric Acid Experiment Isa Free and Paid eBooks

- Metal Carbonate With Hydrochloric Acid Experiment Isa Public Domain eBooks
 - Metal Carbonate With Hydrochloric Acid Experiment Isa eBook Subscription Services
 - Metal Carbonate With Hydrochloric Acid Experiment Isa Budget-Friendly Options
6. Navigating Metal Carbonate With Hydrochloric Acid Experiment Isa eBook Formats
 - ePub, PDF, MOBI, and More
 - Metal Carbonate With Hydrochloric Acid Experiment Isa Compatibility with Devices
 - Metal Carbonate With Hydrochloric Acid Experiment Isa Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Metal Carbonate With Hydrochloric Acid Experiment Isa
 - Highlighting and Note-Taking Metal Carbonate With Hydrochloric Acid Experiment Isa
 - Interactive Elements Metal Carbonate With Hydrochloric Acid Experiment Isa
 8. Staying Engaged with Metal Carbonate With Hydrochloric Acid Experiment Isa
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Metal Carbonate With Hydrochloric Acid Experiment Isa
 9. Balancing eBooks and Physical Books Metal Carbonate With Hydrochloric Acid Experiment Isa
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Metal Carbonate With Hydrochloric Acid Experiment Isa
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Metal Carbonate With Hydrochloric Acid Experiment Isa
 - Setting Reading Goals Metal Carbonate With Hydrochloric Acid Experiment Isa
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Metal Carbonate With Hydrochloric Acid Experiment Isa
 - Fact-Checking eBook Content of Metal Carbonate With Hydrochloric Acid Experiment Isa
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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