






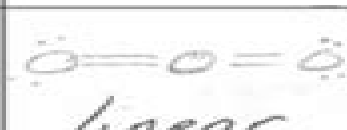
## Molecular Geometry

How can molecular shapes be predicted using the VSEPR theory?

### Why?

When you draw a Lewis structure for a molecule on paper, you are making a two-dimensional representation of the atoms. In reality however, molecules are not flat—they are *three-dimensional*. The true shape of a molecule is important because it determines many physical and chemical properties for the substance. In this activity you will learn how to predict molecular shapes.

### Model 1 – Lewis Structures

Lewis Structures		H <sub>2</sub> CO	3 electron domains (3 bonding, 0 nonbonding)	3-D Molecular Shape
1. H <sub>2</sub> CO				
2. BeF <sub>2</sub>				
3. CH <sub>4</sub>				
4. NH <sub>3</sub>				
5. H <sub>2</sub> O				
6. CO <sub>2</sub>				
Lone pair = ..				

# Geometry Pogil Answer Key

**Thinkwell**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'Thinkwell' text.

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*Geometry Tests Answer Key* Bob Jones University Press,

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