

Chapter 19 Electrochemistry

Practice questions. Answers

1. What element is being reduced in the following redox reaction?



2. What element is being oxidized in the following redox reaction?



3. What element is being oxidized in the following redox reaction?



4. What element is being reduced in the following redox reaction?



5. Balance the following redox reaction if it occurs in acidic solution. What are the coefficients in front of $\text{H}_2\text{C}_2\text{O}_4$ and H_2O in the balanced reaction?



$\text{H}_2\text{C}_2\text{O}_4 = 5, \text{H}_2\text{O} = 8$

6. Balance the following redox reaction if it occurs in acidic solution. What are the coefficients in front of H^+ and Fe^{3+} in the balanced reaction?



$\text{H}^+ = 8, \text{Fe}^{3+} = 5$

7. Determine the cell notation for the redox reaction given below.



$\text{Fe}(\text{s}) \mid \text{Fe}^{3+}(\text{aq}) \parallel \text{Cl}_2(\text{g}) \mid \text{Cl}^-(\text{aq}) \mid \text{Pt}$

8. Use the standard half-cell potentials listed below to calculate the standard cell potential for the following reaction occurring in an electrochemical cell at 25°C . (The equation is balanced.)



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oxidation $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^{-}$ reduction $\text{Fe}^{2+} + 2\text{e}^{-} \rightarrow \text{Fe}$ $24 \text{ 3 Ag} + \text{Ni}^{3+} \rightarrow 3 \text{ Ag} + \text{Ni}$ oxidation reduction $25 \text{ Cu}^{2+} + \text{Pb} \rightarrow \text{Pb}^{2+} + \text{Cu}$ oxidation reduction 26 O_2

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