

Mole and Stoichiometry Mini-Lab

7. Write and balance the equation for the reaction between aluminum and water to produce aluminum oxide and hydrogen gas.



8. Using the moles of the aluminum can in #1 as the GIVEN, how many moles of hydrogen gas can be produced? Assume water is available in excess. UNKNOWN: mol H_2 GIVEN: mol Al (answer to #1)

9. Using the moles of water from #4 as the GIVEN with aluminum in excess, calculate the number of moles of aluminum oxide that can be produced.



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D Keegan



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