

Chapter 19 Practice Problem Set

- Determine whether each of the following process is spontaneous or non-spontaneous.
 - Spontaneous** / nonspontaneous dissolving table salt in hot soup
 - Spontaneous / **nonspontaneous** ball rolling up A Mountain
 - Spontaneous** / nonspontaneous opening a perfume bottle to allow the fragrance to spread through the room
 - Spontaneous / **nonspontaneous** separating helium and neon from a mixture of gases
 - Spontaneous / **nonspontaneous** water solidifying to ice at 3 °C
- Which one of the following processes produces a decrease in the entropy of the system?
 - boiling water to form steam
 - dissolution of solid KCl in water
 - mixing of two gases into one container
 - freezing water to form ice**
 - melting ice to form water
- Which of these species has the largest entropy (S°) at 25°C?
 - CO(g)
 - CH₄(g)**
 - NaCl(s)
 - H₂O(l)
 - Fe(s)
- For each of the following pairs, identify which one has a greater number of microstates (W):
 - CH₄ or **C₂H₆**
 - H₂O(g)** or H₂O(l)
 - 1 mol O₂ occupying 1L at 25°C or **1 mol O₂ occupying 2L at 25°C**
- Predict whether each of the following reactions would have a positive or negative entropy change:
 - $\Delta S > 0$ $S(g) + O_2(g) \rightarrow SO_2(g)$
 - $\Delta S < 0$ $H_2O(l) \rightarrow H_2O(g)$
 - $\Delta S < 0$ $PCl_3(l) + Cl_2(l) \rightarrow PCl_5(l)$
 - $\Delta S > 0$ $2Na(s) + 2H_2O(l) \rightarrow 2NaOH(aq) + H_2(g)$
 - $\Delta S > 0$ $N_2(g) \rightarrow 2N(g)$

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