

- Which event must *always* occur for a chemical reaction to take place?
 - formation of a precipitate
 - formation of a gas
 - effective collisions between reacting particles
 - addition of a catalyst to the reaction system
- In most aqueous reactions as temperature increases, the effectiveness of collisions between reacting particles
 - decreases
 - increases
 - remains the same
 - increases
- A 5.0-gram sample of zinc and a 50.-milliliter sample of hydrochloric acid are used in a chemical reaction. Which combination of these samples has the fastest reaction rate?
 - a zinc strip and 1.0 M HCl(aq)
 - a zinc strip and 3.0 M HCl(aq)
 - zinc powder and 1.0 M HCl(aq)
 - zinc powder and 3.0 M HCl(aq)
- A 1.0-gram piece of zinc reacts with 5 milliliters of HCl(aq). Which of these conditions of concentration and temperature would produce the greatest rate of reaction?
 - 1.0 M HCl(aq) at 20.°C
 - 1.0 M HCl(aq) at 40.°C
 - 2.0 M HCl(aq) at 20.°C
 - 2.0 M HCl(aq) at 40.°C
- At STP, which 4.0-gram zinc sample will react fastest with dilute hydrochloric acid?
 - lump
 - bar
 - powdered
 - sheet metal
- Given the balanced equation representing a reaction:

$$\text{Zn(s)} + 2\text{HCl(aq)} \rightarrow \text{H}_2\text{(g)} + \text{ZnCl}_2\text{(aq)}$$

Which set of reaction conditions produces $\text{H}_2\text{(g)}$ at the fastest rate?

 - a 1.0-g lump of Zn(s) in 50. mL of 0.5 M HCl(aq) at 20.°C
 - a 1.0-g lump of Zn(s) in 50. mL of 0.5 M HCl(aq) at 30.°C
 - 1.0 g of powdered Zn(s) in 50. mL of 1.0 M HCl(aq) at 20.°C
 - 1.0 g of powdered Zn(s) in 50. mL of 1.0 M HCl(aq) at 30.°C
- For a given reaction, adding a catalyst increases the rate of the reaction by
 - providing an alternate reaction pathway that has a higher activation energy
 - providing an alternate reaction pathway that has a lower activation energy
 - using the same reaction pathway and increasing the activation energy
 - using the same reaction pathway and decreasing the activation energy

- Given the balanced equation representing a reaction:



Which statement is true about energy in this reaction?

- The reaction is exothermic because it releases heat.
 - The reaction is exothermic because it absorbs heat.
 - The reaction is endothermic because it releases heat.
 - The reaction is endothermic because it absorbs heat.
- Which balanced equation represents an endothermic reaction?
 - $\text{C(s)} + \text{O}_2\text{(g)} \rightarrow \text{CO}_2\text{(g)}$
 - $\text{CH}_4\text{(g)} + 2\text{O}_2\text{(g)} \rightarrow \text{CO}_2\text{(g)} + 2\text{H}_2\text{O(g)}$
 - $\text{N}_2\text{(g)} + 3\text{H}_2\text{(g)} \rightarrow 2\text{NH}_3\text{(g)}$
 - $\text{N}_2\text{(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{NO(g)}$
 - Given the balanced equation:



Which statement best describes this process?

- It is endothermic and entropy increases.
 - It is endothermic and entropy decreases.
 - It is exothermic and entropy increases.
 - It is exothermic and entropy decreases.
- Which change is exothermic?
 - freezing of water
 - melting of iron
 - vaporization of ethanol
 - sublimation of iodine
 - Which reaction releases the greatest amount of energy per 2 moles of product?
 - $2\text{CO(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{CO}_2\text{(g)}$
 - $4\text{Al(s)} + 3\text{O}_2\text{(g)} \rightarrow 2\text{Al}_2\text{O}_3\text{(s)}$
 - $2\text{H}_2\text{(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{H}_2\text{O(g)}$
 - $\text{N}_2\text{(g)} + 3\text{H}_2\text{(g)} \rightarrow 2\text{NH}_3\text{(g)}$
 - Given the balanced equation representing a reaction at 101.3 kPa and 298 K:

$$\text{N}_2\text{(g)} + 3\text{H}_2\text{(g)} \rightarrow 2\text{NH}_3\text{(g)} + 91.8 \text{ kJ}$$

Which statement is true about this reaction?

 - It is exothermic and ΔH equals -91.8 kJ.
 - It is exothermic and ΔH equals +91.8 kJ.
 - It is endothermic and ΔH equals -91.8 kJ.
 - It is endothermic and ΔH equals +91.8 kJ.
 - Which expression represents the ΔH for a chemical reaction in terms of the potential energy, PE , of its products and reactants?
 - PE of products + PE of reactants
 - PE of products - PE of reactants
 - PE of products $\times$ PE of reactants
 - PE of products $\div$ PE of reactants

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 the set of reactions that may occur among all the species of a chemical system their equilibrium constants and the initial
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