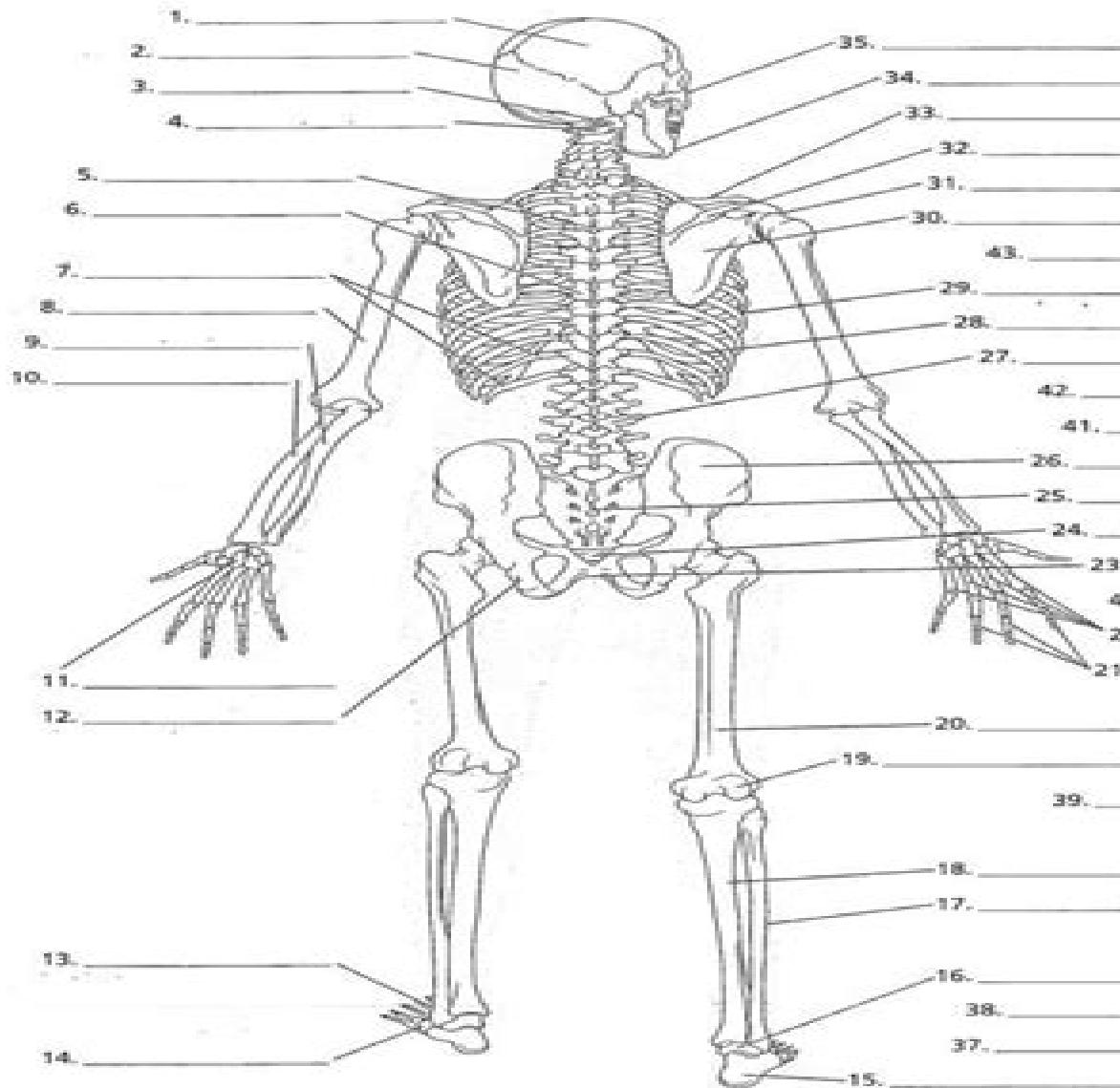
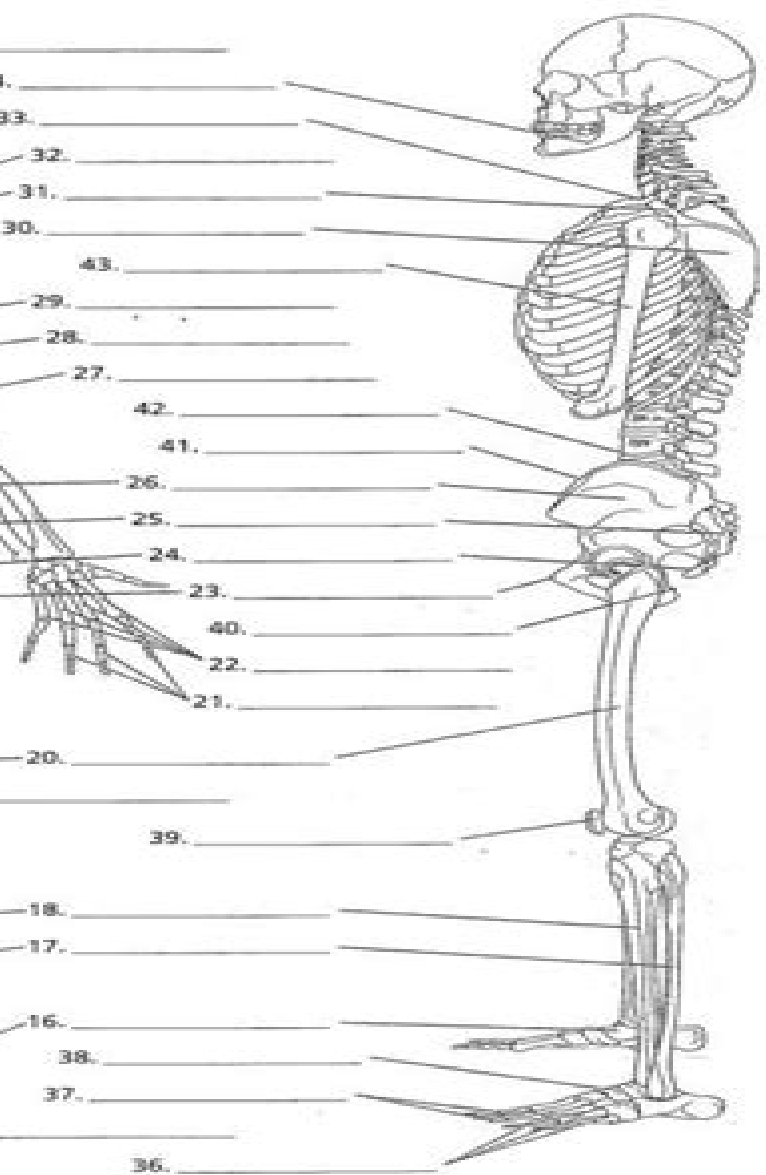


Skeletal System—posterior view



Skeletal System—lateral view



Human Anatomy Physiology Skeletal System Worksheet Answers

R Pring



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subsequent Δh values
 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g}) \quad \Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) + \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \quad \Delta h = 71.2 \text{ kJ}$
 $\text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g}) + \Delta h = 283 \text{ kJ}$

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Hess's law states that if a process can be expressed as the sum of two or more steps the enthalpy change for the overall process is the sum of the ΔH values for each step to use Hess's law two principles must be understood one if an equation is reversed the sign of the ΔH value is also reversed two if an equation is multiplied by a

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$$1. \text{NaOH (s)} \rightarrow \text{Na}^+ \text{aq} + \text{OH}^- \text{aq}$$

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web thus for reaction 9 2 in which 1 mole of H_2SO_4 and 1 mole of $NaOH$ combine to form 2 moles of H_2O $\Delta H = -55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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