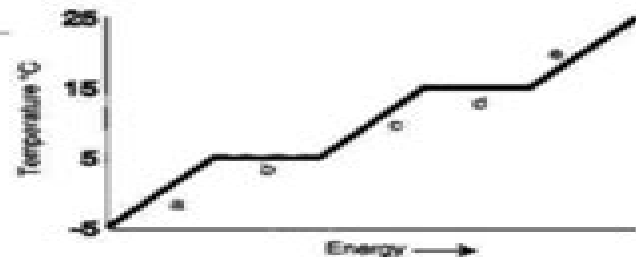


HEATING CURVES QUESTIONS

PART A – HEATING CURVES. Use the heating curve to the right to answer the following questions:

1. What is the melting point of the substance? **B/ 5 C**
2. What is the boiling point of the substance? **D/ 15 C**
3. Which letter represents heating of the solid? **A**
4. Which letter represents heating of the gas? **E**
5. Which letter represents melting of the solid? **B**
6. Which letter represents boiling of the liquid? **D**
7. At points b and d why doesn't the temperature change even though heat energy is added?

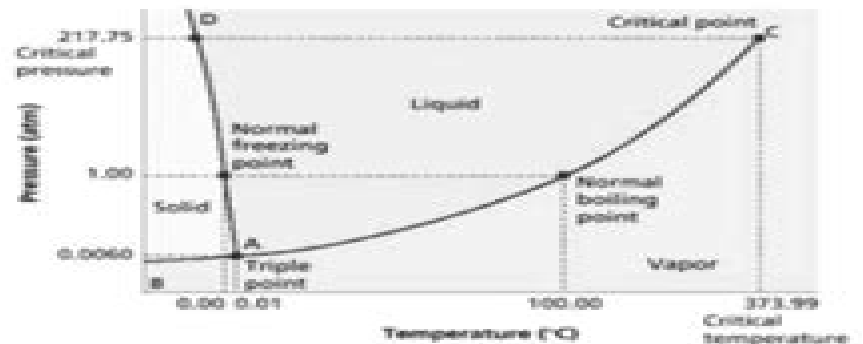


The Substance must absorb enough heat/ kinetic energy to Escape the intermolecular actions/ bonds (i.e. hydrogen bonds) that Keep them their current phase/ state

Refer to the phase diagram below when answering the questions on this worksheet

PART C – PHASE DIAGRAMS. Use the phase diagram for water below to answer the following questions.

8. What is the state of water at 2 atm and 50°? **LIQUID**
9. What phase change will occur if the temperature is lowered from 80°C to -5°C at 1 atm? **FREEZING**
10. You have ice at -10°C and 1 atm. What could you do in order cause the ice to sublime? **Decrease the pressure to less than 0.060 atm**



Predict the physical state of each of these substances at the indicated temperature. Use the melting point and boiling point data from the table below.

- a. phenol at 99°C
- b. ammonia at -25°C
- c. methanol in an ice-water bath
- d. methanol in a boiling-water bath
- e. ammonia at -100°C
- f. phenol at 25°C

- a. **Liquid**
- b. **Gas**
- c. **Liquid**
- d. **Gas**
- e. **Solid**
- f. **Solid**

Heating Cooling Curve Practice Problems 1 Answers

Jung-Chul (Thomas) Eun

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Table of Contents Heating Cooling Curve Practice Problems 1 Answers

1. Understanding the eBook Heating Cooling Curve Practice Problems 1 Answers
 - The Rise of Digital Reading Heating Cooling Curve Practice Problems 1 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Heating Cooling Curve Practice Problems 1 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Heating Cooling Curve Practice Problems 1 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Heating Cooling Curve Practice Problems 1 Answers
 - Personalized Recommendations
 - Heating Cooling Curve Practice Problems 1 Answers User Reviews and Ratings
 - Heating Cooling Curve Practice Problems 1 Answers and Bestseller Lists

5. Accessing Heating Cooling Curve Practice Problems 1 Answers Free and Paid eBooks
 - Heating Cooling Curve Practice Problems 1 Answers Public Domain eBooks
 - Heating Cooling Curve Practice Problems 1 Answers eBook Subscription Services
 - Heating Cooling Curve Practice Problems 1 Answers Budget-Friendly Options
6. Navigating Heating Cooling Curve Practice Problems 1 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Heating Cooling Curve Practice Problems 1 Answers Compatibility with Devices
 - Heating Cooling Curve Practice Problems 1 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Heating Cooling Curve Practice Problems 1 Answers
 - Highlighting and Note-Taking Heating Cooling Curve Practice Problems 1 Answers
 - Interactive Elements Heating Cooling Curve Practice Problems 1 Answers
8. Staying Engaged with Heating Cooling Curve Practice Problems 1 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Heating Cooling Curve Practice Problems 1 Answers
9. Balancing eBooks and Physical Books Heating Cooling Curve Practice Problems 1 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Heating Cooling Curve Practice Problems 1 Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Heating Cooling Curve Practice Problems 1 Answers
 - Setting Reading Goals Heating Cooling Curve Practice Problems 1 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Heating Cooling Curve Practice Problems 1 Answers
 - Fact-Checking eBook Content of Heating Cooling Curve Practice Problems 1 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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