

# Determining the Enthalpy of a Neutralization Reaction Work Sheet

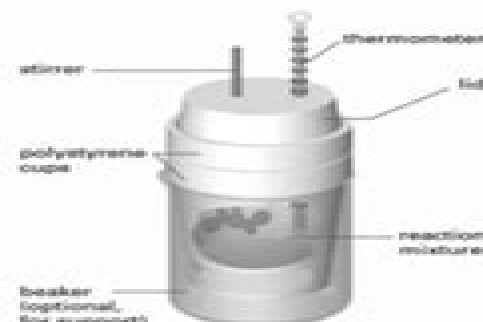
Use the virtual lab.

Follow the procedure on page 240 and 241.

Modifications

- use 1 mol/L HCl, and 3 mol/l HBr.
- use 50 mLs and 100 mLs.
- do three trials each

*Formal lab write up is required as are the following questions. Answer all question using full sentences on a separate sheet of paper.*



1. Explain what is meant by the terms endothermic or exothermic?"
2. Explain what is meant by the term "neutralization reaction"?
3. What is the purpose of the foam cup?
4. Determine the number of moles of HCl(aq) or HBr(aq) and NaOH(aq) that were involved in the reactions. (1 mol and 3 mol 50 mLs and 100 mLs)
5. Calculate the amount of heat that is absorbed by the solution in the calorimeter. List any assumptions that you made as you carried out your calculations. There should be 3.
6. Use the following equation to determine the amount of heat that is released by the reaction:  $-Q_{\text{reaction}} = Q_{\text{solution}}$   
Why does the Q of reaction have an opposite value of the Q of solution?
7. Was heat released or absorbed during the neutralization reaction? Explain your answer.
8. Use your results to determine the enthalpy change if the neutralization reaction, in kJ/mol of NaOH. Write the thermochemical equation for the neutralization reaction.
9. Suppose that you had added solid sodium hydroxide pellets to hydrochloric acid, instead of adding hydrochloric acid to sodium hydroxide solution?
  - a. Would the enthalpy change have been higher or lower?
  - b. How can you test your answer?
  - c. What change do you need to make to the thermochemical equation if you perform the investigation using solid sodium hydroxide?

# Heat Of Neutralization Lab Report

**United States. Environmental  
Protection Agency**



## Heat Of Neutralization Lab Report:

Laboratory Manual for Principles of General Chemistry J. A. Beran, Mark Lassiter, 2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section

**Laboratory Experiments for Brown and LeMay, Chemistry, the Central Science** John Henry Nelson, Kenneth C. Kemp, 1977 *Laboratory Manual to Accompany Chemistry, [by] Stanley R. Radel, Marjorie H. Navidi* Arthur D. Baker, 1990 Fundamentals of Chemistry in the Laboratory Ralph A. Burns, 1999 **Chemical Principles** Ralph J. Christensen, 1998 Laboratory Experiments for General Chemistry Harold R. Hunt, Toby F. Block, 1994 U.S. Government Research & Development Reports, 1969-10 Scientific and Technical Aerospace Reports, 1995 Investigation in General Chemistry Alice S. Cohen, 1978 *Basic Laboratory Principles in General Chemistry* Fitzgerald B. Bramwell, 1990

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