

HEAT AND ITS MEASUREMENT

Name _____

Heat (or energy) can be measured in units of calories or joules. When there is a temperature change (ΔT), heat (Q) can be calculated using this formula:

$$Q = \text{mass} \times \Delta T \times \text{specific heat capacity}$$

($\Delta T = \text{Final Temp} - \text{Initial Temp}$)

During a phase change, we use this formula:

$$Q = \text{mass} \times \text{heat of fusion (or heat of vaporization)}$$

Solve the following problems.

1. How many joules of heat are given off when 5.0 g of water cool from 75° C to 25° C? (Specific heat of water = 4.18 J/g° C)

$$(5)(50)(4.18) = 1045 \text{ rounded } 1000 \text{ J}$$

2. How many calories are given off by the water in Problem 1? (Specific heat of water = 1.0 cal/g° C)

$$(5)(50)(1) = 250 \text{ cal}$$

3. How many joules does it take to melt 35 g of ice at 0° C? (heat of fusion = 333 J/g)

$$(35g)(333) = 11655$$

No ΔT when melting Just 0°

12000 J

4. How many calories are given off when 85 g of steam condense to liquid water? (heat of vaporization = 539.4 cal/g)

$$46000 \text{ cal}$$

5. How many joules of heat are necessary to raise the temperature of 25 g of water from 10° C to 60° C?

$$5000 \text{ J}$$

6. How many calories are given off when 50 g of water at 0° freezes? (heat of fusion = 79.72 cal/g)

$$4000 \text{ cal}$$

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JG Myers



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