Calorimetry: Heat of Neutralization

Prepared by Dongling Fei Manutee Community College

Purpose: Determine the calorimeter constant for a calorimeter. Determine the enthalpy of neutralization for the reaction of a strong acid and a strong base.

Background: When a system absorbs heat (q) its temperature changes by Δ T. Experimental measurements demonstrate that the heat absorbed by a system and its corresponding temperature change are directly proportional. q=C x Δ T. C is called the heat capacity, which is often defined as the quantity of heat required to change its temperature by $1^{\circ}C$

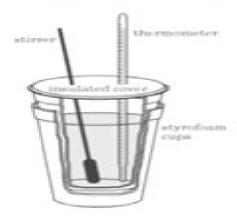
$$C = q/\Delta T = J/C$$

Heat capacity is an extensive property; it depends on the amount of matter being heated. The measure of intrinsic capacity of a substance to absorb heat is called **specific heat capacity** (C_s), the amount of heat required to raise the temperature of 1 gram of the substance by 1 °C

$$C_s=q/(\Delta T *m)$$

The C_s of H₂O is 4.184J/g "C. For the case where all the heat is used to heat water, the amount of heat is calculated by the relation

The enthalpy change for a chemical reaction, ΔH_{ma} is also called the **enthalpy of reaction** or **heat of reaction**, which is also an extensive property. For many aqueous reacions, ΔH_{ma} can be measured simply using a coffee-cup calorimeter



Heat Of Neutralization Lab Manual

Karen Timberlake

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