

# 8 • Electron Configurations & Periodicity

## WRITING ELECTRON CONFIGURATIONS

For each given element, fill in the orbital diagram and then write the electron configuration for the element.

|                                                 |                                                 |                                               |                                                |                                                |                                                 |
|-------------------------------------------------|-------------------------------------------------|-----------------------------------------------|------------------------------------------------|------------------------------------------------|-------------------------------------------------|
| 1.                                              | 2.                                              | 3.                                            | 4.                                             | 5.                                             | 6.                                              |
|                                                 |                                                 |                                               |                                                |                                                |                                                 |
| Element: Ar<br># of e <sup>-</sup> s: <u>18</u> | Element: Mg<br># of e <sup>-</sup> s: <u>12</u> | Element: N<br># of e <sup>-</sup> s: <u>7</u> | Element: Li<br># of e <sup>-</sup> s: <u>3</u> | Element: P<br># of e <sup>-</sup> s: <u>15</u> | Element: Cl<br># of e <sup>-</sup> s: <u>17</u> |

Write the electron configurations of each of these in **long form** and **short form**:

- Ar  $1s^2 2s^2 2p^6 3s^2 3p^6$   
 Ar  $[Ne] 3s^2 3p^6$
- Mg  $1s^2 2s^2 2p^6 3s^2$   
 Mg  $[Ne] 3s^2$
- N  $1s^2 2s^2 2p^3$   
 N  $[He] 2s^2 2p^3$
- Li  $1s^2 2s^1$   
 Li  $[He] 2s^1$
- P  $1s^2 2s^2 2p^6 3s^2 3p^3$   
 P  $[Ne] 3s^2 3p^3$
- Cl  $1s^2 2s^2 2p^6 3s^2 3p^5$   
 Cl  $[Ne] 3s^2 3p^5$



- Fill in the orbital diagram for the element, Fe, and write the electron configuration of Fe in the long and short form.



# Electron Configuration Review Answer Sheet

**Kenneth S. Overway**



## **Electron Configuration Review Answer Sheet:**

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Homework Helpers Chemistry is an excellent review for any standardized Chemistry test including the SAT II its real value is in providing support and guidance during the year's entire course of study      **Foundations of College Chemistry,**

**Alternate** Morris Hein, Susan Arena, 2010-01-26 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years this book has helped them master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They'll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis      *Publications of the National Bureau of Standards ... Catalog* United States. National Bureau of Standards, 1979

Catalog of National Bureau of Standards Publications, 1966-1976: Citations and abstracts United States. National Bureau of Standards. Technical Information and Publications Division, 1978      **NBS Special Publication** , 1968      *Publications of the National Institute of Standards and Technology ... Catalog* National Institute of Standards and Technology (U.S.), 1971

Department of Defense appropriations for fiscal year 1984 United States. Congress. Senate. Committee on Appropriations. Subcommittee on Department of Defense, 1983

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