

7. Electrolysis of water

Electrodes: Platinum or graphite

Electrolyte: dilute sulfuric acid solution

Ions Present: H^+ SO_4^{2-} H^+ OH^-

Half equations:

In the Cathode $(2H^+ + 2e^- \rightarrow H_2) \times 2$

In the Anode $4OH^- \rightarrow O_2 + 2H_2O + 4e^-$

Overall Equation $4H^+ + 4OH^- \rightarrow O_2 + 2H_2 + 2H_2O$

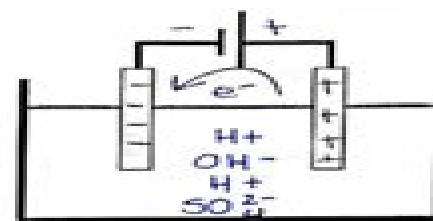
Observations:

Cathode: colorless gas (possibly H_2)

How can you prove it? lighted wood splint pops

Anode: Colorless gas (possibly oxygen)

How can you prove it? Glowing wood splint relights



Label the Hoffman apparatus:

A. oxygen gas

B. anode

C. cathode

D. hydrogen gas

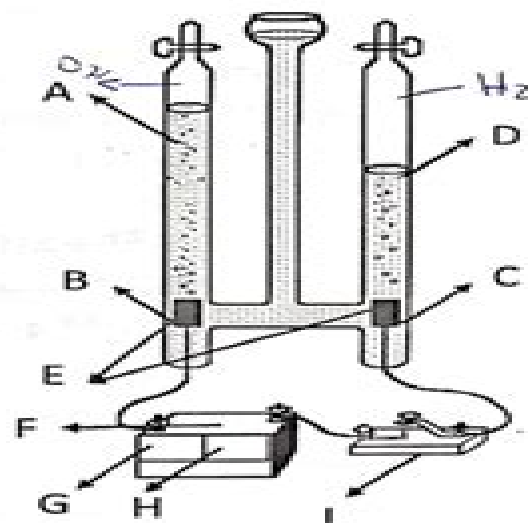
E. electrodes

F. battery

G. (+)

H. (-)

I. switch



Electrochemistry Study Guide Chemistry Matter And Change

Joacim Rocklöv



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