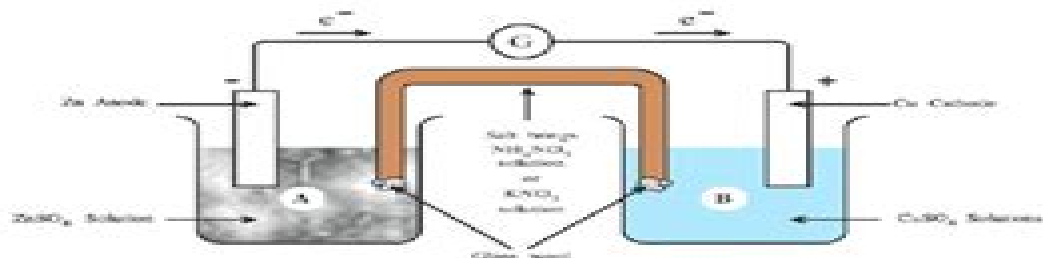


## UNIT : 8 REDOX REACTIONS & ELECTROCHEMISTRY

### Important Points

The device which transforms the chemical energy into electrical energy is called electrochemical cell. In the electrochemical cell, the energy producer in chemical reaction is transformed into electrical energy.



the salt bridge joins the two solutions and keeps the neutrality of electrical charge.

The cell in which the concentration of the ions associated with the reaction of the cell are 1 more litre<sup>-1</sup> and if any gas is associated with the reaction, then its pressure equal to 1 bar and if temperature is 298 K, is called standard cell.

The relative intensity of tendency of receiving the electron is called reduction potential  $E_{red}$  or  $E^{\circ}_{red}$  the relative intensity of electrodes for releasing electron is called oxidation potential

$E_{red}$  or  $E^{\circ}_{red}$

#### Nernst Equation



we can write the concentrations of  $Cu^{2+}$  and  $Zn^{2+}$  ions in Daniell cell, in the given electrode potential.

$$E_{cell} = E^{\circ}_{cell} - \frac{RT}{2F} \ln \left[ \frac{Zn^{2+}}{Cu^{2+}} \right]$$

#### Concentration Cell :

If the two electrodes of the electrochemical cell are same but the Concentration of the solution of the electrolyte are different, then it is called concentration cell. viz



$$\text{Cell reaction: } Cu^{2+}(C_2) \rightleftharpoons Cu^{2+}(C_1) \quad E_{cell} = E^{\circ}_{cell} - \frac{0.059}{n} \log \left[ \frac{C_1}{C_2} \right]$$

# Electrochemistry Practical Manual

**Brendan G. Carr**



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