

$$\Delta S [\text{H}_2\text{O}_2 (g)] = 232.9 \text{ J K}^{-1} \text{ mol}^{-1} \quad \Delta H_f [\text{H}_2\text{O}_2 (g)] = -136.1 \text{ kJ mol}^{-1}$$

$$\Delta S_{\text{universe}} = \Delta S_{\text{sys}} + \Delta S_{\text{surr}} = 5.3 \text{ J K}^{-1} \text{ mol}^{-1} \text{ spontaneous}$$

7. The reaction  $\text{NH}_3(g) + \text{HCl}(g) \rightarrow \text{NH}_4\text{Cl}(s)$  is spontaneous at room temperature. Determine the temperature at which it is no longer spontaneous. Given information

$$S^\circ [\text{NH}_3(g)] = 193.0 \text{ J K}^{-1} \text{ mol}^{-1}$$

$$S^\circ [\text{HCl}(g)] = 187.0 \text{ J K}^{-1} \text{ mol}^{-1}$$

$$S^\circ [\text{NH}_4\text{Cl}(s)] = 94.56 \text{ J K}^{-1} \text{ mol}^{-1}$$

$$\Delta H_f^\circ [\text{NH}_3(g)] = -46.3 \text{ kJ mol}^{-1}$$

$$\Delta H_f^\circ [\text{HCl}(g)] = -92.3 \text{ kJ mol}^{-1}$$

$$\Delta H_f^\circ [\text{NH}_4\text{Cl}(s)] = -315.39 \text{ kJ mol}^{-1}$$

$$T < 619 \text{ K}$$

8. Calculate the change in the entropies of the system and the surroundings, and the total change in entropy, when the volume of a sample of argon gas of mass 21 g at 298 K and 1.50 bar increases from 1.20 dm<sup>3</sup> to 4.60 dm<sup>3</sup> in (a) an isothermal reversible expansion, (b) an isothermal irreversible expansion against  $p_{\text{ex}} = 0$ , and (c) an adiabatic reversible expansion

$$\text{a) } \Delta S_{\text{system}} = +3.0 \text{ J/K}, \Delta S_{\text{surrounding}} = -3.0 \text{ J/K}, \Delta S_{\text{total}} = 0 \text{ J/K}$$

$$\text{b) } \Delta S_{\text{system}} = +3.0 \text{ J/K}, \Delta S_{\text{surrounding}} = -3.0 \text{ J/K}, \Delta S_{\text{total}} = +6.0 \text{ J/K}$$

$$\text{c) } \Delta S_{\text{system}} = +3.0 \text{ J/K}, \Delta S_{\text{surrounding}} = -3.0 \text{ J/K}, \Delta S_{\text{total}} = +6.0 \text{ J/K}$$

9. The standard molar enthalpy of formation of  $\text{H}_2\text{O}(g)$  is  $-241.8 \text{ kJ mol}^{-1}$  and its  $\Delta G_f^\circ$  is  $-228.6 \text{ kJ mol}^{-1}$ .

Calculate the standard molar entropies of  $\text{H}_2(g)$  and  $\text{O}_2(g)$  at 298 K, using the following equation and the following equation in Table 2.2. Calculate the standard molar entropies of  $\text{H}_2(g)$  and  $\text{O}_2(g)$  at 298 K.

$$\Delta G_f^\circ = -228.6 \text{ kJ mol}^{-1}$$

$$\Delta G_f^\circ = -228.6 \text{ kJ mol}^{-1}$$

10. The standard molar enthalpy of formation of  $\text{H}_2\text{O}(g)$  is  $-241.8 \text{ kJ mol}^{-1}$  and its  $\Delta G_f^\circ$  is  $-228.6 \text{ kJ mol}^{-1}$ . Calculate the standard molar entropies of  $\text{H}_2(g)$  and  $\text{O}_2(g)$  at 298 K, using the following equation and the following equation in Table 2.2. Calculate the standard molar entropies of  $\text{H}_2(g)$  and  $\text{O}_2(g)$  at 298 K.

$$\Delta G_f^\circ = -228.6 \text{ kJ mol}^{-1}$$

11. The standard molar enthalpy of formation of  $\text{H}_2\text{O}(g)$  is  $-241.8 \text{ kJ mol}^{-1}$  and its  $\Delta G_f^\circ$  is  $-228.6 \text{ kJ mol}^{-1}$ . Calculate the standard molar entropies of  $\text{H}_2(g)$  and  $\text{O}_2(g)$  at 298 K.

# Exercises For Thermodynamics Answers

**F Rizvi**



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undergraduates The manual is free to all adopters of the main text *Student's Solutions Manual to Accompany Atkins' Physical Chemistry* C. A. Trapp, M. P. Cady, Carmen Giunta, 2010 This solutions manual provides the authors detailed solutions to exercises and problems in physical chemistry It comprises solutions to exercises at the end of each chapter and solutions to numerical theoretical and additional problems

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*Physical Chemistry Student Solutions Manual* Charles Trapp, Peter Atkins, Marshall Cady, 2006-08-11 Change 21

*Thermodynamics* Patrick Jacobs, 2013-03-14 This textbook on thermodynamics is intended primarily for honours and B Sc students majoring in physical chemistry However students of physics engineering and biochemistry will also find the book relevant to their studies Its principal features are a much shorter presentation of the laws of thermodynamics than is customary made possible by the definition of the thermodynamic scale of temperature using only one fixed point the triple point of water which immediately follows the Zeroth Law The author's first exposure to thermodynamics revealed that its usefulness seemed to be mostly confined to the study of gases in equilibrium Readers of this book will find that applications of thermodynamics to liquids and solids as well as gases are emphasized and they will learn that thermodynamics can be applied to systems which are not in equilibrium This book contains three learning aids Fully worked out examples are included at appropriate places in the text which also includes numerous exercises These are designed to help the reader stop and think about what he or she has just read Answers to the exercises are given at the end of each section and there are also problems at the end of each chapter which readers can work out on their own a

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