Name:

Honors Chem Test: Ch. 7

	1	How many atoms of fluorine are present in a molecule of carbon tetrafluoride, CFV			
		a. 1	e.	4	
		b. 2	d.	5	
	2.	Name the compound Al ₂ S ₃			
		a. aluminum sulfate	- C	aluminum(II) sulfate	
		 aluminum sulfur 	d.	aluminum sulfide	
	3.	Name the compound CF.			
		a. cal crum fluoride	c.	carbon tetrafiuon de	
		 b. carbon fluoride 	d.	monocalcium quadrafluoride	
	4.	Name the compound N_2O_3 .			
		a. dimitrogen oonde	C	mitric coulde	
		b. nitrogen triconde	d.	dinitrogen triconde	
_	5.	What is the formula for deastsenic trisulfide?			
		a. Ar ₂ S ₃	· c	As ₂ S ₄	
		b. ArSı	d.	A53S3	
	6.	What is the formula for sulfur dichloride?			
		a. SCI	0.00	S ₂ CI	
		b. SCI;	d.	S,CI,	
	7.	What is the formula for hydrochloric ac	id?		
		a. HF	, c.	HCIO	
		ь нсі	d.	H ₂ CO ₂	
	8.	The molar mass of an element is the mass of one			
		a. atom of the element.	c.	gram of the element.	
		 b. crystal of the element. 	d.	mole of the element.	
-	9.	The molar mass of MgI ₂ is			
		a. the sum of the masses of 1 mol of Mg and 2 mol of 1.			
		 the sum of the masses of 1 mol of Mg and 1 mol of I. 			
		c. the sum of the masses of 1 atom of			

Honors Chemistry Chapter Practice Test

John T. Moore, Richard H. Langley

Honors Chemistry Chapter Practice Test:

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- 1. The book is structured into several chapters, namely:
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 - Chapter 3: Honors Chemistry Chapter Practice Test in Everyday Life
 - Chapter 4: Honors Chemistry Chapter Practice Test in Specific Contexts
 - ∘ Chapter 5: Conclusion
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- 3. In chapter 2, the author will delve into the foundational concepts of Honors Chemistry Chapter Practice Test. The second chapter will elucidate the essential principles that need to be understood to grasp Honors Chemistry Chapter Practice Test in its entirety.
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- 6. In chapter 5, this book will draw a conclusion about Honors Chemistry Chapter Practice Test. This chapter will summarize the key points that have been discussed throughout the book.
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