

Question 2 : Extended Questions**(20 Marks)**

Write your answer for each of the following questions in the space provided.

Be sure to show all your work, including the correct units where applicable.

1- Vinegar is used as a food preservative. It is an acidic solution containing ethanoic acid, CH_3COOH . (a) A titration was carried out to measure the concentration of ethanoic acid in a sample of vinegar. 25.0 cm^3 of a vinegar solution was titrated with a solution of sodium hydroxide, concentration $0.250 \text{ mol dm}^{-3}$. The concentration of the ethanoic acid in the vinegar solution was found to be $0.125 \text{ mol dm}^{-3}$.

(i) Calculate the pH of $0.250 \text{ mol dm}^{-3}$ sodium hydroxide at 298 K .

[$K_w = 1.00 \times 10^{-14} \text{ mol}^2 \text{ dm}^{-3}$ at 298 K .]

(3marks)

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(ii) Define Bronsted acid?

(2marks)

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(iii) Write an equation to show the ionization of the Bronsted -lowery base NH_3 in water (2 marks)

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Victor M. Corman



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