

$$M_{\text{AgNO}_3} = \frac{m_{\text{AgNO}_3}}{V_{\text{AgNO}_3}} = \frac{25.513 \text{ g} + 149.87 \frac{\text{g}}{100 \text{ mL}}}{100.00 \text{ mL}} \times \frac{1000 \text{ mL}}{1 \text{ L}} = \boxed{1.569 \frac{\text{mol}}{\text{L}}}$$

4. Confirm that the molarity is the same molarity that you calculated in step 3.

Yes, the molarity is the same as the molarity calculated in step 3.

### Part B: Unknown solutions

Using a method similar to the one used in part A, create and carry out a procedure for determining the molarities of the unknown  $\text{AgNO}_3$  solution and the unknown  $\text{NaCl}$  solution. Round your data and show your calculations.

#### Unknown $\text{AgNO}_3$

Upon the addition of solid  $\text{NaCl}$ , solid  $\text{AgCl}$  was formed.

$\text{AgCl}$  formed (g) = 2.32019 g

$$\text{mass}(\text{AgNO}_3) = \frac{M_r(\text{AgNO}_3)}{M_r(\text{AgCl})} \times \text{mass}(\text{AgCl}) = \frac{149.87}{143.32} \times 2.32019 \text{ g} = \boxed{2.37500 \text{ g}}$$

$$M_{\text{AgNO}_3} = \frac{m_{\text{AgNO}_3}}{V_{\text{AgNO}_3}} = \frac{2.37500 \text{ g} + 149.87 \frac{\text{g}}{100 \text{ mL}}}{100.00 \text{ mL}} \times \frac{1000 \text{ mL}}{1 \text{ L}} = \boxed{0.16689 \frac{\text{mol}}{\text{L}}}$$

#### Unknown $\text{NaCl}$

Upon the addition of solid  $\text{AgNO}_3$ , solid  $\text{AgCl}$  was formed.

$\text{AgCl}$  formed (g) = 10.7907 g

$$\text{mass}(\text{NaCl}) = \frac{M_r(\text{NaCl})}{M_r(\text{AgCl})} \times \text{mass}(\text{AgCl}) = \frac{58.443}{143.32} \times 10.7907 \text{ g} = \boxed{4.4002 \text{ g}}$$

$$M_{\text{NaCl}} = \frac{m_{\text{NaCl}}}{V_{\text{NaCl}}} = \frac{4.4002 \text{ g} + 58.443 \frac{\text{g}}{100 \text{ mL}}}{100.00 \text{ mL}} \times \frac{1000 \text{ mL}}{1 \text{ L}} = \boxed{0.75294 \frac{\text{mol}}{\text{L}}}$$

### Conclusion questions

1. If you have a company that is using this method to create  $\text{AgCl}$ , knowing that  $\text{AgNO}_3$  is an expensive chemical, what do you think would be a better limiting reagent to use,  $\text{AgNO}_3$  or  $\text{NaCl}$ ? Why?

# Experiment Limiting Reactant Answers

**James E. Brady, Fred Senese**



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## **Table of Contents Experiment Limiting Reactant Answers**

1. Understanding the eBook Experiment Limiting Reactant Answers
  - The Rise of Digital Reading Experiment Limiting Reactant Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Experiment Limiting Reactant Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Experiment Limiting Reactant Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Experiment Limiting Reactant Answers
  - Personalized Recommendations
  - Experiment Limiting Reactant Answers User Reviews and Ratings
  - Experiment Limiting Reactant Answers and Bestseller Lists
5. Accessing Experiment Limiting Reactant Answers Free and Paid eBooks
  - Experiment Limiting Reactant Answers Public Domain eBooks

- Experiment Limiting Reactant Answers eBook Subscription Services
- Experiment Limiting Reactant Answers Budget-Friendly Options
- 6. Navigating Experiment Limiting Reactant Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Experiment Limiting Reactant Answers Compatibility with Devices
  - Experiment Limiting Reactant Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Experiment Limiting Reactant Answers
  - Highlighting and Note-Taking Experiment Limiting Reactant Answers
  - Interactive Elements Experiment Limiting Reactant Answers
- 8. Staying Engaged with Experiment Limiting Reactant Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Experiment Limiting Reactant Answers
- 9. Balancing eBooks and Physical Books Experiment Limiting Reactant Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Experiment Limiting Reactant Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Experiment Limiting Reactant Answers
  - Setting Reading Goals Experiment Limiting Reactant Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Experiment Limiting Reactant Answers
  - Fact-Checking eBook Content of Experiment Limiting Reactant Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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