

Chemical Reactions

Section 9.1 Reactions and Equations

pages 282–288

Practice Problems

pages 284–287

Write skeleton equations for the following word equations.

1. Hydrogen and bromine gases react to yield hydrogen bromide.

hydrogen(g) + bromine(g) →
hydrogen bromide(g)



2. When carbon monoxide and oxygen react, carbon dioxide forms.

carbon monoxide(g) + oxygen(g) →
carbon dioxide(g)



3. **Challenge** Write the word equation and the skeleton equation for the following reaction: when heated, solid potassium chlorate yields solid potassium chloride and oxygen gas.

Potassium chlorate(s) →
potassium chloride(s) + oxygen(g)



Write chemical equations for each of the following reactions.

4. In water, iron(III) chloride reacts with sodium hydroxide, producing solid iron(III) hydroxide and sodium chloride.



5. Liquid carbon disulfide reacts with oxygen gas, producing carbon dioxide gas and sulfur dioxide gas.



6. **Challenge** A piece of zinc metal is added to a solution of hydrogen sulfate. This reaction produces a gas and a solution of zinc sulfate.



Section 9.1 Assessment

page 288

7. **Explain** why it is important that a chemical equation be balanced.

Because mass is neither created nor destroyed in chemical reactions, the numbers of atoms of all elements must be equal on both sides of the reaction arrow.

8. **List** three types of physical evidence that indicate a chemical reaction has occurred.

Answers may include release or absorption of energy, change in color, change in odor, formation of a gas, or formation of a solid.

9. **Compare and contrast** a skeleton equation and a chemical equation.

The skeleton equation includes the formulas of reactants and products. The chemical equation gives the relative amounts of reactants and products.

10. **Explain** why it is important to reduce coefficients in a balanced equation to the lowest possible whole-number ratio.

Coefficients in the lowest ratio most clearly indicate the relative amounts of substances in a reaction.

11. **Analyze** When balancing a chemical equation, can you adjust the subscript in a formula? Explain.

No. Doing so changes the identity of the substance.

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Julia Schneider



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Table of Contents Mcgraw Hill Chemistry Chap 2 Assessment Answers

1. Understanding the eBook Mcgraw Hill Chemistry Chap 2 Assessment Answers
 - The Rise of Digital Reading Mcgraw Hill Chemistry Chap 2 Assessment Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Mcgraw Hill Chemistry Chap 2 Assessment Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mcgraw Hill Chemistry Chap 2 Assessment Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mcgraw Hill Chemistry Chap 2 Assessment Answers

- Personalized Recommendations
 - Mcgraw Hill Chemistry Chap 2 Assessment Answers User Reviews and Ratings
 - Mcgraw Hill Chemistry Chap 2 Assessment Answers and Bestseller Lists
5. Accessing Mcgraw Hill Chemistry Chap 2 Assessment Answers Free and Paid eBooks
 - Mcgraw Hill Chemistry Chap 2 Assessment Answers Public Domain eBooks
 - Mcgraw Hill Chemistry Chap 2 Assessment Answers eBook Subscription Services
 - Mcgraw Hill Chemistry Chap 2 Assessment Answers Budget-Friendly Options
 6. Navigating Mcgraw Hill Chemistry Chap 2 Assessment Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Mcgraw Hill Chemistry Chap 2 Assessment Answers Compatibility with Devices
 - Mcgraw Hill Chemistry Chap 2 Assessment Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mcgraw Hill Chemistry Chap 2 Assessment Answers
 - Highlighting and Note-Taking Mcgraw Hill Chemistry Chap 2 Assessment Answers
 - Interactive Elements Mcgraw Hill Chemistry Chap 2 Assessment Answers
 8. Staying Engaged with Mcgraw Hill Chemistry Chap 2 Assessment Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mcgraw Hill Chemistry Chap 2 Assessment Answers
 9. Balancing eBooks and Physical Books Mcgraw Hill Chemistry Chap 2 Assessment Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mcgraw Hill Chemistry Chap 2 Assessment Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Mcgraw Hill Chemistry Chap 2 Assessment Answers
 - Setting Reading Goals Mcgraw Hill Chemistry Chap 2 Assessment Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Mcgraw Hill Chemistry Chap 2 Assessment Answers

- Fact-Checking eBook Content of Mcgraw Hill Chemistry Chap 2 Assessment Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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