

9 • Bonding & Molecular Structure

PRACTICE TEST

1. The correct Lewis symbol for ground state carbon



2. The correct Lewis symbol for ground state aluminum is



3. Using the picture below, what process corresponds to the lattice energy?



4. Which of the following favors formation of an ionic compound?

- (a) low ionization energy for metal
(b) high electron affinity for non metal
(c) high lattice energy
(d) all of a-c above
e) none of a-c above

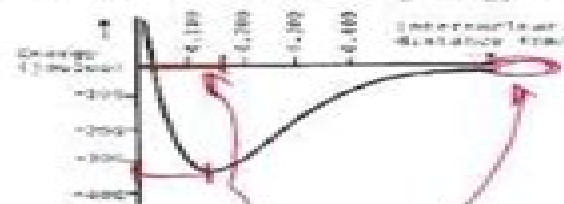
5. Which of the atoms below is least likely to violate the octet rule?



6. How many electrons are shown in the Lewis structure of perchlorate ion, ClO_4^- ?



- Questions 7 - 9 refer to the following energy diagram:



7. What is the energy of the two isolated atoms?

- a) -400 J d) 0.100 nm
b) -335 J e) -155 J
(c) 0 J

8. What is the bond length of the bond between the two atoms?

- a) 0.020 nm d) 0.330 nm
(b) 0.140 nm e) 2.00 nm
c) 0.400 nm

9. What is the bond energy (bond strength) of the bond?

- a) 0 J (d) -330 J Should this be + value? Unknown
b) -110 J e) -422 J
c) -10 J

10. What is the bond order of the C-O bond in acetone?



- a) 4 b) 1.5 c) 0.5 d) 1 (c) 2

double = B.O 2

Mcdougal Geometry Chapter2 Test Answer Key

Ron Larson, McDougal Littell



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