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shape of the molecule drawing and name polar or nonpolar p or np h 2 o

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web the pcl 3 lewis structure has the typical case of phosphorus p in the center with 3 bonds to 3 other atoms phosphorus is

from the same column as nitrogen in the periodic table meaning that p and n generally have the same bonding structure note the lone pair dots without bonds on top of p just like for n in the previous example for NH_3 chlorine cl

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web step 1 figure out how many electrons the molecule must have based on the number of valence electrons in each atom when drawing the structure of an ion be sure to add subtract electrons to account for the charge step 2 connect the atoms to each other with single bonds to form a skeleton structure

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