Hyann I Study Claide

- Physical Changes vs. Chemical Changes (Can you undo 17).
 - Physical changes result in changes of appearance, but not composition.
 - Melting, evaporation, freezing, etc.
 - Chemical changes result in changes of composition of matter (Chemical reaction)
 - Burning, mixing, dissolving, etc.
- Physical Properties vs. Chemical Properties
 - Physical properties can be described by observing the matter without a change in the composition.
 - · Color, odor, density, temperature, etc.
 - Chemical properties can only be detected by aftering the composition.
 - Flammability, incidination, etc.
- Anoma va Molecules.
 - Atoms are the smallest unit of a given element that carry the properties of that element.
 - Molecules are composed of two or more atoms.
- Elements vs. Compounds
 - Elements occur in nature and cannot be broken down to simpler particles through ordinary means.
 - Compounds are the combination of two or more elements.
- Accuracy vs. Precision
 - Accuracy refers to the agreement of a measurement with a known or accepted value. (Bullseyer)
 - Correctioness.
 - Procision refers to the agroement of a measurement with other measurements in a given set of measurements. (Wissing the bullsaye but landing all darts in one place)
 - Repeatability or reproducibility
- Rules Concerning Significant Figures
 - All nomore digits ARE significant.
 - Zeroes between nonzero digits ARE significant.
 - Zeroes at the beginning of a number are NOT significant, they only indicate the position of the documal point.
 - Zeros at the end of a number to the right of a decimal point ARE significant.
 - Zeroes at the end of a number with no decimal point are ASSUMED to be NOT significant, UNLESS we indicate their significance by placing a line over the last zero that is significant.

General Chemistry Standardized Final Exam Study Guide

Sabine Zange

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