

# 7 Ionic and Metallic Bonding

## **Big idea** BONDING AND INTERACTIONS

### 7.1 Ions

For students using the Foundation edition, assign problems 1, 3–5, 7–12, 14, 15, 18–20.

**Essential Understanding** Ions form when atoms gain or lose valence electrons, becoming electrically charged.

#### Lesson Summary

**Valence Electrons** Valence electrons are the electrons in the outermost occupied energy level and are involved in ion formation.

- ▶ For a representative element, the group number equals the number of valence electrons the atom contains.
- ▶ An electron dot structure shows the symbol of the element and its valence electrons.
- ▶ Atoms tend to gain or lose the number of electrons that will provide the atom with a noble gas electron configuration.

**Formation of Cations** Cations are positively charged ions formed when an atom loses one or more valence electrons.

- ▶ Atoms and the cations formed from them have different properties.
- ▶ Elements in Group 1 form cations with a charge of  $1+$ , and those in Group 2 form cations with a charge of  $2+$ .
- ▶ Many transition metals form more than one cation and do not follow the octet rule.

**Formation of Anions** Anions are negatively charged ions formed when an atom gains one or more valence electrons.

- ▶ Commonly, the name of an anion ends in *-ide*.
- ▶ Anions form from nonmetallic elements.
- ▶ The anions formed from halogens are known as halides.

After reading Lesson 7.1, answer the following questions.

#### Valence Electrons

1. What are valence electrons?

*Valence electrons are the electrons in the highest occupied energy level of an element's atoms.*

2. The valence electrons largely determine the *chemical properties* of an element and are usually the only electrons used in *chemical bonds*.

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