

# Pass the Proton Acid-Base Theories

Name \_\_\_\_\_

Date \_\_\_\_\_ Period \_\_\_\_\_

**Purpose**

To compare acid and base solutions at the particulate level.

**Materials**

- Acid-Base Solution cards

**Procedure and Questions**

- Sort the cards into three groups based on common features. Write the solution card number, molarity, chemical formula, and chemical name for each solution.

## Group 1

4: CHOOH  
1: HCL  
5: CH<sub>2</sub>COOH  
6: NaCl  
9: NaOH  
7: CH<sub>2</sub>OH  
11: NH<sub>4</sub>

All have 0.010 M

## Group 2

2: HCL  
8: CH<sub>2</sub>OH

All have 0.005 M

## Group 3

3: HCL  
10: NaOH  
12: HOH

All have 0.001 M

- What do the solutions in each of the three groups have in common?

Group 1: same molarity, 0.010 M and a lot of protons

Group 2: same molarity, 0.005M and little protons

Group 3: same molarity, 0.001 M and lack of protons

- Is there a second way you might sort the cards? If so, explain how.

You can sort them by the patterns/shape of protons

- Fill in the table with the chemical formula of the substance dissolved and the ions in the solution.

# Lesson 12 Unit 4 Toxins Worksheet

**S Ashworth**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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