Same Class Dute

Assessment

# Chapter Test B

# Atomic Physics

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

What is the frequency of a photon with an energy of 1.99 × 10<sup>-19</sup> J?

$$(h = 6.63 \times 10^{-14} \text{ Jes})$$

a. 
$$1.00 \times 10^{14} \, \mathrm{Hz}$$

d. 
$$4.00 \times 10^{14} \, \mathrm{Hz}$$

2. Light with an energy equal to three times the work function of a given metal causes the metal to eject photoelectrons. What is the ratio of the maximum photoelectron kinetic energy to the work function?

3. A monochromatic light beam with a quantum energy value of 3.0 eV is incident upon a photocell. The work function of the photocell is 1.6 eV. What is the maximum kinetic energy of the ejected electrons?

- 4. What causes the bright lines in the emission spectrum of an element to occur?
  - a. Photons are absorbed when electrons jump from a higher-energy to a lower-energy state.
  - b. Photons are emitted when electrons jump from a higher-energy to a lower-energy state.
  - c. Photons are absorbed when electrons jump from a lower-energy to a higher-energy state.
  - d. Photons are emitted when electrons jump from a lower-energy to a higher-energy state.
  - \_\_\_ 5. What causes the dark lines in the absorption spectrum of an element to occur?
    - a. Photons are absorbed when electrons jump from a higher-energy to a lower-energy state.
    - b. Photons are emitted when electrons jump from a higher-energy to a lower-energy state.
    - c. Photons are absorbed when electrons jump from a lower-energy to a higher-energy state.
    - d. Photons are emitted when electrons jump from a lower-energy to a higher-energy state.

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**Marco Cascella** 

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