

100) Carbon tetrachloride consists of one carbon and four chlorine atoms. How many chlorine atoms are present in 1.00 mole of CCl_4 molecules?

$$\text{one mole } \text{CCl}_4 \text{ contains } 4 \text{ moles of } \text{Cl} \quad \text{one mole } \text{CCl}_4 \text{ contains } 4 \text{ moles of } \text{Cl}$$

$$1 \text{ mole } \text{CCl}_4 \text{ contains } 4 \text{ moles of } \text{Cl}$$

$$\text{answer is } 4 \text{ moles of } \text{Cl} \quad \text{answer is } 4$$

101) There are 4 monovalent atoms in each sulfur hexafluoride molecule.

102) Describe how each of the following is formed:

1. Distillation separation of mixtures

2. Crystallization separation of mixtures

3. Extraction separation of mixtures

103) Suggest other three non ionic separation techniques in solid state.

1. Distillation separation of mixtures

2. Crystallization separation of mixtures

104) Sodium chromate is a compound of sulfur, oxygen, sodium, and chromium.

105) Calculate the number of atoms in the formula $\text{Na}_2\text{Cr}_2\text{O}_7$.

What is the oxidation number of sulfur in $\text{Na}_2\text{Cr}_2\text{O}_7$?



F321 Chemistry May 2013

Mike Smith



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