



### Rates of Reaction Worksheet

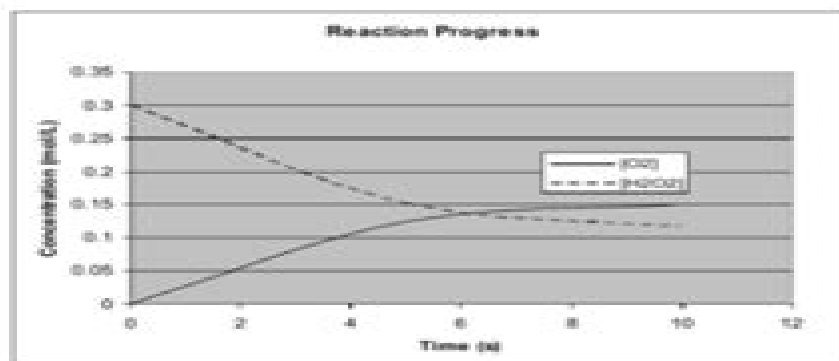
1. What are the five most common ways to determine reaction rates?
2. Solid phosphorus and oxygen gas react to form tetraphosphorus decoxide. Determine the average rate of reaction for oxygen during the first 40 s if the concentration changes from 0.200 mol/L to 0.000 mol/L during this time interval.
3. At 40°C, hydrogen chloride gas will form from the reaction of gaseous hydrogen and chlorine, according to the following balanced chemical equation:  $\text{H}_2(\text{g}) + \text{Cl}_2(\text{g}) \rightarrow 2 \text{HCl}(\text{g})$ . Using the data provided, calculate the following average rates:

Time (s)	Concentration (mol/L)		
	$\text{H}_2(\text{g})$	$\text{Cl}_2(\text{g})$	$\text{HCl}(\text{g})$
0	1.000	1.000	0.000
2.16	0.500	0.500	1.000
4.32	0.250	0.250	1.500

- a) disappearance of hydrogen gas in the first 2.16 s. (0.231 mol/L.s)
- b) appearance of hydrogen chloride gas in the first 4.32 s (0.347 mol/L.s)
- c) disappearance of chlorine gas between 2.16 s and 4.32 s (0.116 mol/L.s)
- d) disappearance of hydrogen gas in the first 4.32 s (0.174 mol/L.s)

4. Hydrogen peroxide in aqueous solution will decompose to produce oxygen gas and water. Use the graph to the right to:

- a) calculate the average reaction rate for the appearance of oxygen gas between 5 s and 10 s.
- b) calculate the average reaction rate of disappearance of hydrogen peroxide between 0 s and 5 s.
- c) calculate the instantaneous rate of appearance of oxygen gas at 4 s.



5. Ammonia gas and oxygen gas react to form nitric oxide and water vapour. The instantaneous rate of consumption of ammonia is  $2.0 \times 10^{-2}$  mol/L at 10 s.
  - a) Write and balance the chemical equation.
  - b) Determine the instantaneous rate of oxygen gas consumption at 10 s. ( $2.2 \times 10^{-2}$  mol/L.s)
  - c) Determine the instantaneous rate of water vapour appearance at 10 s. ( $2.7 \times 10^{-2}$  mol/L.s)
6. The chemical reaction represented by the following balanced chemical equation takes place in aqueous solution:  $\text{MnO}_4^- (\text{aq}) + 5 \text{Fe}^{2+} (\text{aq}) + 8 \text{H}^+ (\text{aq}) \rightarrow \text{Mn}^{2+} (\text{aq}) + 5 \text{Fe}^{3+} (\text{aq}) + 4 \text{H}_2\text{O} (\text{l})$ . When the disappearance of the permanganate ion,  $\text{MnO}_4^- (\text{aq})$ , is monitored, the rate of reaction for the consumption of the ion is  $4.0 \times 10^{-2}$  mol/L.s. Calculate the reaction rate for all other reactants and products. ( $\text{Fe}^{2+}$ : 0.20 mol/L.s,  $\text{H}^+$ : 0.32 mol/L.s,  $\text{Mn}^{2+}$ : 0.040 mol/L.s,  $\text{Fe}^{3+}$ : 0.20 mol/L.s,  $\text{H}_2\text{O}$ : 0.16 mol/L.s)
7. One mole of aqueous sulfuric acid reacts with 2 mol of aqueous sodium hydrogen carbonate in an acid-base reaction.
  - a) Write a balanced chemical equation.
  - b) If 0.038 mol sodium hydrogen carbonate reacts in 20 s, calculate the average reaction rate as:
    - i. Grams of sodium hydrogen carbonate consumed per second. (0.16 g/s)
    - ii. Moles of sulfuric acid consumed per second. ( $9.5 \times 10^{-4}$  mol/s)
    - iii. Moles of carbon dioxide gas formed per second. ( $1.9 \times 10^{-3}$  mol/s)

## Lesson 36 Reaction Rates Answer

**Kathy Moscou, Karen Snipe**



### Lesson 36 Reaction Rates Answer:

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