

Empirical/Molecular Formula

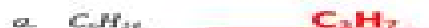
Name: _____

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1. What is an empirical formula? How is it different from a molecular formula?

The empirical formula is the lowest whole number ratio of atoms in a compound. The molecular formula shows the true number of atoms in the molecule and is a whole number multiple of the empirical formula.

2. State the empirical formula for each of the following.



3. A sample of a compound is 50.0% sulfur and 50.0% oxygen by mass. What is the empirical formula?

$$S: \frac{50.0 \text{ g S}}{32.07 \text{ g S}} \times \frac{1 \text{ mol S}}{1} = \frac{1.559 \text{ mol S}}{1.559 \text{ mol}} = 1 \text{ S}$$

EF: SO_2

$$O: \frac{50.0 \text{ g O}}{16.00 \text{ g O}} \times \frac{1 \text{ mol O}}{1} = \frac{3.125 \text{ mol O}}{1.559 \text{ mol}} = 2 \text{ O}$$

4. A sample of a compound is 52.9% aluminum. The rest is oxygen. What is the empirical formula?

$$Al: \frac{52.9 \text{ g Al}}{26.98 \text{ g Al}} \times \frac{1 \text{ mol Al}}{1} = \frac{1.961 \text{ mol Al}}{1.961 \text{ mol Al}} = 1 \text{ Al} \times 2 = 2$$

EF: Al_2O_3

$$O: \frac{47.1 \text{ g O}}{16.00 \text{ g O}} \times \frac{1 \text{ mol O}}{1} = \frac{2.944 \text{ mol O}}{1.961 \text{ mol}} = 1.5 \times 2 = 3 \text{ O}$$

Ratio must be a whole

5. A compound is 48.65% carbon, 8.11 % hydrogen and 43.24 % oxygen. Determine the empirical formula.

4.05 mol carbon, 8.11 mol H, 2.70 mol O



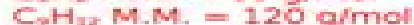
1. Divide by the smallest # to get = $C_{1.5}H_3O_1$

2. Multiply to get a whole number ratio (x2) = $C_3H_6O_2$

6. Vinegar is a dilute solution of acetic acid. The molar mass of acetic acid is 60.0 g/mol and it has an empirical formula of CH_2O . What is the molecular formula of acetic acid?



7. A compound has an empirical formula of C_3H_4 . Which of the following are possible molar masses of the compound? 20 g/mol, 55 g/mol, 80 g/mol, 120 g/mol.



Molecular Formula Answers

K Payea

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Table of Contents Molecular Formula Answers

1. Understanding the eBook Molecular Formula Answers
 - The Rise of Digital Reading Molecular Formula Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Molecular Formula Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Molecular Formula Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Molecular Formula Answers
 - Personalized Recommendations
 - Molecular Formula Answers User Reviews and Ratings
 - Molecular Formula Answers and Bestseller Lists
5. Accessing Molecular Formula Answers Free and Paid eBooks
 - Molecular Formula Answers Public Domain eBooks
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 - Molecular Formula Answers Budget-Friendly Options
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- Molecular Formula Answers Enhanced eBook Features
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 - Highlighting and Note-Taking Molecular Formula Answers
 - Interactive Elements Molecular Formula Answers
- 8. Staying Engaged with Molecular Formula Answers
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 - Participating in Virtual Book Clubs
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 - Minimizing Distractions
 - Managing Screen Time
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 - Setting Reading Goals Molecular Formula Answers
 - Carving Out Dedicated Reading Time
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 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
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