

## MOLE RATIOS IN CHEMICAL EQUATIONS

When the equation is balanced, you can use the **coefficients** to get the **mole ratios** between all substances present to get the **amount** (moles) of each (pure) substance.

Remember, that one mole of a substance is equal to Avogadro's number ( $N_A$ ) =  $6.022 \times 10^{23}$  particles/mol of **particles** or **entities** of that substance (which could be a reactant or product). Below is a list of what an 'entity' represents for each type of pure substance.

**Element** – atom e.g. Al

**Covalent/Molecular Compound** – molecule e.g.  $H_2SO_4$

**Ionic Compound** – formula unit e.g.  $Al_2(SO_4)_3$

There is more than one way to set up/look at a mole ratio, but ultimately the final answer should be the same. There are short cuts too, but you should only use these after you understand what is happening.



### MOLE RATIOS



so, if you have 5.00 moles of Al, then how many moles of  $H_2SO_4$  are required to react with the Al?

*"cross multiplication" method*

$$\begin{aligned} 2 \text{ Al} &: 3 \text{ H}_2\text{SO}_4 \\ \frac{2 \text{ Al}}{5.00 \text{ Al}} &= \frac{3 \text{ H}_2\text{SO}_4}{x \text{ mol H}_2\text{SO}_4} \\ &= \frac{3 \text{ H}_2\text{SO}_4 \times 5.00 \text{ Al}}{2 \text{ Al}} \end{aligned}$$

*G x R/G method*

$$\begin{aligned} G \times \frac{R}{G} &= 5.00 \text{ Al} \times \frac{3 \text{ H}_2\text{SO}_4}{2 \text{ Al}} \\ &= 7.50 \text{ mols H}_2\text{SO}_4 \end{aligned}$$

↑  
(coefficients)

# Mole Ratios Model 1 A Chemical Reaction Answers

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Jiménez**



## **Mole Ratios Model 1 A Chemical Reaction Answers:**

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