











at How do the reduced ratios in the last column compare to the coefficients in the reaction shown in Model 1?

The calm are the same at the coefficients in stabil 1 - 108-18.

A. Use mathematical concepts to explain how your answer in part a is possible.

The coefficients were all published by the come months; so they at other hast to the same thing-

- Even 538 is a small number of molecules to use in a reaction. Typically chemists use much larger numbers of molecules. (Recall that one mole is equal to 6.02 x 10<sup>th</sup> particles.) Consider each situation below as it relates to the reaction in Model 1: N<sub>c</sub>(g) + 3H<sub>c</sub>(g) → 2NH<sub>c</sub>(g).
  - at. Calculate the amount of reactants consumed and products made.
  - A. Record the ratio of N, to H, to NH, Reduce the ratio to the lowest whole number possible.

	N <sub>2</sub> Consumed	H <sub>1</sub> Consumed	NH, Produced	Ratio N <sub>1</sub> H <sub>2</sub> NH <sub>3</sub>
If the reaction occurred 6.02 × 10 <sup>23</sup> times, how many molecules would be consumed or pro- duced?	6.01 * 15 <sup>th</sup>	1.61 - 10**	120-10"	1:3:2
How many moles of each sub- stance would be consumed or produced in the previous situa- tion?	1	3	2	1:3:2

- 5. Refer to the data table in Question 4.
  - a. How do the reduced ratios in the last column corepare to the coefficients in the reaction in Model 12

The reduced palms are the same at the coefficients.

A. Use mathematical concepts to explain how your answer in part a is possible.

ble are all authorizing coefficients by the same make acque just their energiating by their

- 6. The ratio obtained from the coefficients in a balanced chemical equation is called the mole ratio.
  - a. What is the mole ratio for the reaction in Model 1?

7 7 2

6. Explain why this ratio is called the mole ratio?

It is the reduced ratio is unles (all the conficients a fine with





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**L** Towne



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