

Balance the equation below



27) Once the equation is balanced, what is the **molar ratio** of water (H₂O) to oxygen (O₂)? 1:1

28) If 4.59 grams of O₂ react, how many grams of CH₄ are needed?

$$\frac{4.59 \text{ g-O}_2}{1} \times \frac{1 \text{ mole-O}_2}{32.00 \text{ g-O}_2} \times \frac{1 \text{ mole-CH}_4}{2 \text{ moles-O}_2} \times \frac{16.05 \text{ g CH}_4}{1 \text{ mole-CH}_4}$$

Answer: 1.15 g CH₄

29) How many grams of CO₂ are needed to react with 8.95 grams of H₂O?

$$\frac{8.95 \text{ g-H}_2\text{O}}{1} \times \frac{1 \text{ mole-H}_2\text{O}}{18.02 \text{ g-H}_2\text{O}} \times \frac{1 \text{ mole-CO}_2}{2 \text{ moles-H}_2\text{O}} \times \frac{44.01 \text{ g CO}_2}{1 \text{ mole-CO}_2}$$

Answer: 10.93 g CO₂

Balance the equation below



30) Once the equation is balanced, what is the **molar ratio** of H₂ to O₂? 2:1

31) How many moles of H₂ are needed to react with 0.42 moles of H₂O?

$$\frac{0.42 \text{ moles-H}_2\text{O}}{1} \times \frac{2 \text{ moles H}_2}{2 \text{ moles-H}_2\text{O}}$$

Answer: 0.42 moles H₂

32) How many grams of H₂O are needed to react with 4.55 moles of H₂?

$$\frac{4.55 \text{ moles-H}_2}{1} \times \frac{2 \text{ moles-H}_2\text{O}}{2 \text{ moles-H}_2} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mole-H}_2\text{O}}$$

Answer: 81.99 g H₂O

33) How many moles of H₂O are produced when reacted with 8.35 grams of H₂?

$$\frac{8.35 \text{ g-H}_2}{1} \times \frac{1 \text{ mole-H}_2}{2.02 \text{ g-H}_2} \times \frac{2 \text{ moles H}_2\text{O}}{2 \text{ moles-H}_2}$$

Answer: 4.13 moles H₂O

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S Marginson



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