

10Lab | Experiment 5: The Mole, Counting by Weighing

Last Name _____ First _____

____ / ____ pts

Postlab Questions

Part 4. Calculations. Converting from Mass to Moles for some common objects.

Using the mass of the iron cube, the 50.0mL of water and the packet of sugar, determine the number of moles of metal, water and sugar. Round of to correct number of significant figures and use scientific notation when appropriate.

J The number of moles of water in 50.0 ml of water

	Part4 J
J1) 50.0 mL of water: Write the mass of the 50-ml water from part J.	1) _____ g 50.0 mL water
J2) Write the molar mass of water.	2) _____ g/mole for water
J3) Calculate the number of moles of water in the 50-mL volume of water. Show your calculations below.	3) _____ moles of water
J4) Calculate the number of water molecules in 50.0 mL. Show your calculations below.	4) _____ molecules of water
J5) Calculate the number of Oxygen and H atoms in 50 mL of water. Show your calculations below.	5a) _____ O atoms in water
	5b) _____ H atoms in water

J The number of moles of sucrose in a packet of sugar

J6) Sucrose (sugar), $C_{12}H_{22}O_{11}$: Write the mass of sucrose.	6) _____ g mass of sucrose.
J7) Write the molar mass of sucrose	7) _____ g/mol for sucrose.
J8) Calculate the number of moles of the sucrose. Show your calculation below.	8) _____ moles of sucrose
J9) Calculate the number of molecules of sucrose in the packet of sugar. Show your calculations below.	9) _____ molecules of sugar
J10) Calculate the number of atoms of C, H, and O in the packet of sucrose. Show your calculations below.	10a) _____ C atoms
	10b) _____ H atoms
	10c) _____ O atoms

Mole Lab Counting And Weighing Answers

Neil D. Jespersen, Alison Hyslop



Mole Lab Counting And Weighing Answers:

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Chemistry in the Community. American Chemical Society,2002 This volume has relevance to a wide number of courses giving a hands on introduction to chemistry in relation to community issues rather than around specific chemical concepts

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