

Mole Calculation Practice Worksheet

Answer the following questions:

- 1) How many moles are in 25 grams of water?

$$\frac{25 \text{ g H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{1 \text{ mol H}_2\text{O}}{1} = 1.4 \text{ moles}$$

- 2) How many grams are in 4.5 moles of Li_2O ?

$$\frac{4.5 \text{ mol Li}_2\text{O}}{1 \text{ mol Li}_2\text{O}} \times \frac{29.88 \text{ g Li}_2\text{O}}{1} = 130 \text{ g}$$

- 3) How many molecules are in 23 moles of oxygen?

$$\frac{23 \text{ mol O}_2}{1 \text{ mol O}_2} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1} = 1.4 \times 10^{25} \text{ molecules}$$

- 4) How many moles are in 3.4×10^{23} molecules of H_2SO_4 ?

$$\frac{3.4 \times 10^{23} \text{ molecules H}_2\text{SO}_4}{6.02 \times 10^{23} \text{ molecules H}_2\text{SO}_4} \times \frac{1 \text{ mol H}_2\text{SO}_4}{1} = 0.56 \text{ moles}$$

- 5) How many molecules are in 25 grams of NH_3 ?

$$\frac{25 \text{ g NH}_3}{17.03 \text{ g NH}_3} \times \frac{1 \text{ mol NH}_3}{1} \times \frac{6.02 \times 10^{23} \text{ molecules NH}_3}{1 \text{ mol NH}_3} = 8.8 \times 10^{23} \text{ molecules}$$

- 6) How many grams are in 8.2×10^{22} molecules of N_2I_4 ?

$$\frac{8.2 \times 10^{22} \text{ molecules N}_2\text{I}_4}{6.02 \times 10^{23} \text{ molecules N}_2\text{I}_4} \times \frac{1 \text{ mol N}_2\text{I}_4}{1} \times \frac{789.42 \text{ g N}_2\text{I}_4}{1 \text{ mol N}_2\text{I}_4} = 110 \text{ g}$$

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L Darling-Hammond



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