

# Molarity Practice problems

$$M = \frac{\text{mol}}{L}$$



Key

1. How many moles of  $\text{CaCl}_2$  would be dissolved in 1.0 L of a 0.1 M solution of  $\text{CaCl}_2$ ?

$$\text{mol} = M \times L = 0.1 \frac{\text{mol}}{L} \times 1 L = 0.1 \text{ mol } \text{CaCl}_2$$

2. What is the concentration of 25 mL of solution that has 0.05 moles of solute dissolved in it?

$$M = \frac{\text{mol}}{L} = \frac{0.05 \text{ mol}}{0.025 L} = 2 M$$

3. How many liters of water need to be added to 0.5 moles of  $\text{CaCl}_2$  to make a 0.20 M solution of  $\text{CaCl}_2$ ?

$$L = \frac{\text{mol}}{M} = \frac{0.5 \text{ mol}}{0.20 M} = 2.5 L$$

4. How many grams of potassium carbonate are needed to make 200 mL of a 2.5 M solution?

$$\text{mol} = M \times L \quad \text{K}_2\text{CO}_3$$

$$\text{mol} = 2.5 M \times 0.200 L = 0.50 \text{ mol}$$

$$\frac{0.50 \text{ mol } \text{K}_2\text{CO}_3}{1 \text{ mol}} \times 138.21 \text{ g} = 69.1 \text{ g } \text{K}_2\text{CO}_3$$

5. How many grams of Lithium Bromide would be needed to make 1.43 L of 4 M Lithium Bromide solution?

$$\text{mol} = M \times L \quad \text{LiBr}$$

$$= 4 M \times 1.43 L$$

$$= 5.72 \text{ mol } \text{LiBr}$$

$$\frac{5.72 \text{ mol } \text{LiBr}}{1 \text{ mol}} \times 86.84 \text{ g} = 496.7 \text{ g } \text{LiBr}$$

6. What is the concentration of a 450 mL solution that contains 200 grams of iron (II) chloride?  $\text{FeCl}_2$

$$M = \frac{\text{mol}}{L}$$

$$\frac{200 \text{ g } \text{FeCl}_2}{126.75 \text{ g}} \times \frac{1 \text{ mol}}{1 \text{ mol}} = 1.578 \text{ mol } \text{FeCl}_2$$

$$\frac{1.578 \text{ mol}}{0.450 L} = 3.5 M$$

7. How many grams of ammonium sulfate are needed to make a 0.25 L solution at a concentration of 6 M?

$$\text{mol} = M \times L \quad (\text{NH}_4)_2\text{SO}_4$$

$$= 6 M \times 0.25 L$$

$$= 1.5 \text{ mol } (\text{NH}_4)_2\text{SO}_4$$

$$\frac{1.5 \text{ mol } (\text{NH}_4)_2\text{SO}_4}{1 \text{ mol}} \times 132.138 \text{ g} = 198.2 \text{ g } (\text{NH}_4)_2\text{SO}_4$$

8. What is the concentration of a solution that has a volume of 2.5 L and contains 660 grams of calcium phosphate?

$$M = \frac{\text{mol}}{L}$$

$$\frac{660 \text{ g } \text{Ca}_3(\text{PO}_4)_2}{310.18 \text{ g}} \times \frac{1 \text{ mol}}{1 \text{ mol}} = 2.127 \text{ mol}$$

$$\frac{2.127 \text{ mol}}{2.5 L} = 0.85 M$$

9. How many grams of copper (II) fluoride are needed to make 6.7 liters of a 1.2 M solution?

$$\text{mol} = M \times L \quad \text{CuF}_2$$

$$= 1.2 M \times 6.7 L$$

$$= 8.04 \text{ mol } \text{CuF}_2$$

$$\frac{8.04 \text{ mol } \text{CuF}_2}{1 \text{ mol}} \times 101.55 \text{ g} = 816.5 \text{ g } \text{CuF}_2$$

# Molarity Practice Problems Answer Key With Work

**CL Gary**



**Molarity Practice Problems Answer Key With Work:**

## Unveiling the Power of Verbal Art: An Psychological Sojourn through **Molarity Practice Problems Answer Key With Work**

In a global inundated with monitors and the cacophony of instantaneous interaction, the profound power and mental resonance of verbal art usually fade in to obscurity, eclipsed by the regular onslaught of noise and distractions. Yet, situated within the musical pages of **Molarity Practice Problems Answer Key With Work**, a interesting perform of literary elegance that pulses with fresh feelings, lies an memorable journey waiting to be embarked upon. Penned by way of a virtuoso wordsmith, this mesmerizing opus instructions visitors on a psychological odyssey, softly exposing the latent potential and profound impact embedded within the delicate internet of language. Within the heart-wrenching expanse with this evocative evaluation, we shall embark upon an introspective exploration of the book is main themes, dissect their interesting publishing design, and immerse ourselves in the indelible effect it leaves upon the depths of readers souls.

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