

Example:

25. Three grams of carbon and 15 g of oxygen are placed in a closed vessel and heated to 1000°C. What mass of carbon dioxide is formed?

$$25 \text{ g } C + \frac{1 \text{ mol } C}{12 \text{ g } C} = \frac{8}{12} \text{ mol } C = \frac{2}{3} \text{ mol } C$$

$$15 \text{ g } O_2 = \frac{1 \text{ mol } O_2}{32 \text{ g } O_2} = \frac{1}{2} \text{ mol } O_2$$

$$\frac{2}{3} \text{ mol } C + \frac{1}{2} \text{ mol } O_2 = \frac{4}{6} \text{ mol } CO_2 = \frac{2}{3} \text{ mol } CO_2$$

$$\frac{2}{3} \text{ mol } CO_2 \times \frac{44 \text{ g } CO_2}{1 \text{ mol } CO_2} = \boxed{29.33 \text{ g } CO_2}$$

26. When 10 g of aluminum reacts with oxygen, 25 g of  $Al_2O_3$  is formed. How much oxygen is required to form 25 g of  $Al_2O_3$ ?



$$10 \text{ g } Al = \frac{1 \text{ mol } Al}{27 \text{ g } Al} = \frac{1}{2.7} \text{ mol } Al$$

$$\frac{1}{2.7} \text{ mol } Al \times \frac{3 \text{ mol } O_2}{4 \text{ mol } Al} = \frac{3}{10.8} \text{ mol } O_2$$

$$\frac{3}{10.8} \text{ mol } O_2 \times \frac{32 \text{ g } O_2}{1 \text{ mol } O_2} = 9.33 \text{ g } O_2$$

$$\% \text{ yield} = \frac{25 \text{ g } Al_2O_3}{9.33 \text{ g } O_2} \times 100 = \boxed{267.8\%}$$

27. In the reaction of problem 26, what is the limiting reagent when 10 g of Al and 15 g of  $O_2$  are used?

$$10 \text{ g } Al = \frac{1 \text{ mol } Al}{27 \text{ g } Al} = \frac{1}{2.7} \text{ mol } Al$$

$$15 \text{ g } O_2 = \frac{1 \text{ mol } O_2}{32 \text{ g } O_2} = \frac{1}{2.13} \text{ mol } O_2$$

$$\frac{1}{2.7} \text{ mol } Al \times \frac{3 \text{ mol } O_2}{4 \text{ mol } Al} = \frac{3}{10.8} \text{ mol } O_2$$

$$\frac{3}{10.8} \text{ mol } O_2 \times \frac{32 \text{ g } O_2}{1 \text{ mol } O_2} = 9.33 \text{ g } O_2$$

Limiting  
Reagent = Al

# Module 8 Stoichiometry Answer Key

**Camilla Rothe**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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### **Table of Contents Module 8 Stoichiometry Answer Key**

1. Understanding the eBook Module 8 Stoichiometry Answer Key
  - The Rise of Digital Reading Module 8 Stoichiometry Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Module 8 Stoichiometry Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Module 8 Stoichiometry Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Module 8 Stoichiometry Answer Key
  - Personalized Recommendations
  - Module 8 Stoichiometry Answer Key User Reviews and Ratings

- Module 8 Stoichiometry Answer Key and Bestseller Lists
- 5. Accessing Module 8 Stoichiometry Answer Key Free and Paid eBooks
  - Module 8 Stoichiometry Answer Key Public Domain eBooks
  - Module 8 Stoichiometry Answer Key eBook Subscription Services
  - Module 8 Stoichiometry Answer Key Budget-Friendly Options
- 6. Navigating Module 8 Stoichiometry Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Module 8 Stoichiometry Answer Key Compatibility with Devices
  - Module 8 Stoichiometry Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Module 8 Stoichiometry Answer Key
  - Highlighting and Note-Taking Module 8 Stoichiometry Answer Key
  - Interactive Elements Module 8 Stoichiometry Answer Key
- 8. Staying Engaged with Module 8 Stoichiometry Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Module 8 Stoichiometry Answer Key
- 9. Balancing eBooks and Physical Books Module 8 Stoichiometry Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Module 8 Stoichiometry Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Module 8 Stoichiometry Answer Key
  - Setting Reading Goals Module 8 Stoichiometry Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Module 8 Stoichiometry Answer Key
  - Fact-Checking eBook Content of Module 8 Stoichiometry Answer Key
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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