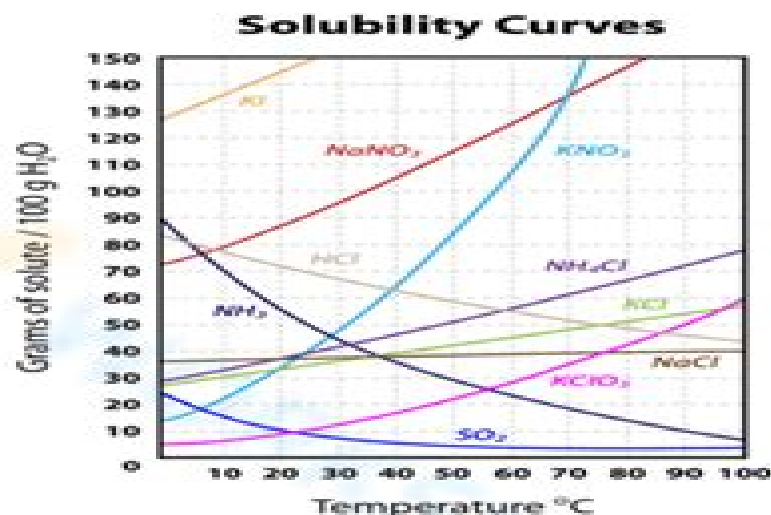


# WS 6.1: Interpreting Solubility Curves

*Directions: Use the Solubility Curve Table below to answer all questions. So you do not have to keep flipping your paper over you might want to put the solubility curve on your phone or computer screen when answering the questions on the back. All questions assume the amount of water is 100 mL.*



- What is the solvent in this Solubility Curve?  
**water**
- Saturation points are given per how much solvent?  
**100 ml or 100 g**
- What is the saturation point of NaNO<sub>3</sub> at 10°C, 40°C, & 80°C?  
 @ 10°C = ~79 g  
 @ 40°C = ~105 g  
 @ 80°C = ~146 g
- How much NH<sub>3</sub> can you dissolve in water at 10°C, 30°C, & 90°C?  
 @ 10°C = ~70 g  
 @ 30°C = ~45 g  
 @ 90°C = ~10 g
- What is the maximum amount of KCl that can be dissolved in water at 20°C, 50°C, & 90°C?  
 @ 20°C = ~33 g  
 @ 50°C = ~42 g  
 @ 90°C = ~53 g
- If you want to make a saturated solution of NaCl how many grams would you dissolve in 100 grams of water @ 80°C?  
**About 40 grams**
- Except for KI What substance is the most soluble in water at 20°C?  
**NaNO<sub>3</sub>**
- Except for KI, what substance is the most soluble in water at 72°C?  
**KNO<sub>3</sub>**
- Look up The boiling points for SO<sub>2</sub>, HCl and NH<sub>3</sub> and write them below:  
 B.P. SO<sub>2</sub> = -10°C  
 B.P. HCl = -85°C  
 B.P. NH<sub>3</sub> = -33°C
- Based on the boiling points you looked up, are SO<sub>2</sub>, HCl & NH<sub>3</sub> solids, liquids, or gases at 0°C? (and anywhere else on the solubility curve)  
**gases**

# Interpreting Solubility Curves Answers

**Richard Bailey**



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## **Table of Contents Interpreting Solubility Curves Answers**

1. Understanding the eBook Interpreting Solubility Curves Answers
  - The Rise of Digital Reading Interpreting Solubility Curves Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Interpreting Solubility Curves Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Interpreting Solubility Curves Answers
  - User-Friendly Interface

4. Exploring eBook Recommendations from Interpreting Solubility Curves Answers
  - Personalized Recommendations
  - Interpreting Solubility Curves Answers User Reviews and Ratings
  - Interpreting Solubility Curves Answers and Bestseller Lists
5. Accessing Interpreting Solubility Curves Answers Free and Paid eBooks
  - Interpreting Solubility Curves Answers Public Domain eBooks
  - Interpreting Solubility Curves Answers eBook Subscription Services
  - Interpreting Solubility Curves Answers Budget-Friendly Options
6. Navigating Interpreting Solubility Curves Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Interpreting Solubility Curves Answers Compatibility with Devices
  - Interpreting Solubility Curves Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Interpreting Solubility Curves Answers
  - Highlighting and Note-Taking Interpreting Solubility Curves Answers
  - Interactive Elements Interpreting Solubility Curves Answers
8. Staying Engaged with Interpreting Solubility Curves Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Interpreting Solubility Curves Answers
9. Balancing eBooks and Physical Books Interpreting Solubility Curves Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Interpreting Solubility Curves Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Interpreting Solubility Curves Answers
  - Setting Reading Goals Interpreting Solubility Curves Answers
  - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Interpreting Solubility Curves Answers
  - Fact-Checking eBook Content of Interpreting Solubility Curves Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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