

Graphing Exercise #1

Richard C. Evers, University of Wisconsin - (Cold gas) Equations
© 2009, Dr. Evers

Equation of state form:

$$P/P^* = Z(T)$$

For this equation for the real gas, we need a few assumptions about the behavior of the gas. We will use the van der Waals equation of state:

1. The gas is a pure substance.
2. The gas is a pure substance.
3. The gas is a pure substance.

There are three other assumptions we will use. There are three other assumptions we will use for the van der Waals equation of state. There are three other assumptions we will use for the van der Waals equation of state.

$$\left[P + \left(\frac{a}{V^2} \right) \right] (V - b) = RT$$

The two terms that are subtracted are a/V^2 and b . The term a/V^2 is the attractive force between molecules. The term b is the volume of the molecules.

The term a/V^2 is the attractive force between molecules. The term b is the volume of the molecules. The term a/V^2 is the attractive force between molecules. The term b is the volume of the molecules.

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