

1. Consider the following electron configurations:

i)	$1s^2 2s^2 3s^2$
ii)	$1s^2 2s^2 2p^6$
iii)	$s^2 2s^2 2p^6 3s^1$
iv)	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^8$
v)	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6$

Which configurations represent elements in the same group on the periodic table?  
 A) ii) and v)      B) i) and iv)      C) iii) and v)      D) ii) and iii)

ANS: C

1. What is the electron configuration of the sulfide ion,  $S^{2-}$ ?

- A)  $1s^2 2s^2 2p^6 3s^2 3p^4$     B)  $1s^2 2s^2 2p^6 3s^2 3p^6$   
 C)  $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2$     D)  $1s^2 2s^2 2p^6 3s^2 3p^3$

ANS: B

1.  $[Ar]4s^2 3d^7 4p^1$  is the electron configuration of which \_\_\_\_\_ atom.

- A) As      B) V      C) P      D) Sb

ANS: A

1. There are \_\_\_\_\_ unpaired electrons in a ground state phosphorus atom.

- A) 0      B) 1      C) 2      D) 3      E) 4

ANS: D

1. There are \_\_\_\_\_ unpaired electrons in a ground state fluorine atom.

- A) 0      B) 1      C) 2      D) 3      E) 4

ANS: B

1. In a ground-state manganese atom, the \_\_\_\_\_ subshell is partially filled.

- A) 3s      B) 4s      C) 4p      D) 3d      E) 4d

ANS: D

1. The electron configuration of a ground-state Ag atom is \_\_\_\_\_.

- A)  $[Ar]4s^2 4d^9$       B)  $[Kr]5s^2 4d^9$       C)  $[Kr]4s^2 4d^9$       D)  $[Ar]4s^2 4d^9$       E)  $[Kr]5s^2 4d^9$

ANS: B

1. The ground state electron configuration for Zn is \_\_\_\_\_.

# Electron Configuration Multiple Choice Questions

**RD Boyd**



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## **Table of Contents Electron Configuration Multiple Choice Questions**

1. Understanding the eBook Electron Configuration Multiple Choice Questions
  - The Rise of Digital Reading Electron Configuration Multiple Choice Questions
  - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configuration Multiple Choice Questions
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Electron Configuration Multiple Choice Questions
  - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configuration Multiple Choice Questions
  - Personalized Recommendations
  - Electron Configuration Multiple Choice Questions User Reviews and Ratings
  - Electron Configuration Multiple Choice Questions and Bestseller Lists

5. Accessing Electron Configuration Multiple Choice Questions Free and Paid eBooks
  - Electron Configuration Multiple Choice Questions Public Domain eBooks
  - Electron Configuration Multiple Choice Questions eBook Subscription Services
  - Electron Configuration Multiple Choice Questions Budget-Friendly Options
6. Navigating Electron Configuration Multiple Choice Questions eBook Formats
  - ePub, PDF, MOBI, and More
  - Electron Configuration Multiple Choice Questions Compatibility with Devices
  - Electron Configuration Multiple Choice Questions Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Electron Configuration Multiple Choice Questions
  - Highlighting and Note-Taking Electron Configuration Multiple Choice Questions
  - Interactive Elements Electron Configuration Multiple Choice Questions
8. Staying Engaged with Electron Configuration Multiple Choice Questions
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Electron Configuration Multiple Choice Questions
9. Balancing eBooks and Physical Books Electron Configuration Multiple Choice Questions
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Electron Configuration Multiple Choice Questions
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Electron Configuration Multiple Choice Questions
  - Setting Reading Goals Electron Configuration Multiple Choice Questions
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Electron Configuration Multiple Choice Questions
  - Fact-Checking eBook Content of Electron Configuration Multiple Choice Questions
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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