#### Additional Exercises: Galvanic Cell

1. Fill in the table below regarding the redox process in a galvanic cell.

Redox Process	Definition in terms of electron movement	Occurs at which half-cell
Oxidation	Loss of electron	Anode
Reduction	Gain of electron	Cathode

- 2. For each of the following pair of electrodes, determine the:
  - cathode and anode
  - half-reactions at the cathode and anode
  - cell notation
  - the standard cell voltage
    - a) Cu & Co

Cathode: Cu, reaction:  $Cu^{2*}(aq) + 2e \rightarrow Cu(s)$ Anode: Co, reaction:  $Co(s) \rightarrow Co^{2*}(aq) + 2e$ Cell notation:  $Co(s) | Co^{2*}(aq) | | Cu^{2*}(aq) | Cu(s)$  $E^0 = 0.34V - (-0.28V) = 0.62V$ 

b) Zn & Ag

Cathode: Ag, reaction: Ag'(aq) + e  $\rightarrow$  Ag(s) Anode: Zn, reaction: Zn(s)  $\rightarrow$  Zn<sup>2\*</sup>(aq) +2e Cell notation: Zn(s) | Zn<sup>2\*</sup>(aq) | Ag'(aq) | Ag(s)  $\mathcal{E}^{\bullet}$  = 0.80V -(-0.76V) = 1.56V

c) Mg & Al

Cathode: Al, reaction:  $Al^{2*}(aq) + 3e \rightarrow Al(s)$ Anode: Mg, reaction: Mg(s)  $\rightarrow$  Mg<sup>2\*</sup>(aq) +2e Cell notation: Mg(s) [Mg<sup>2\*</sup>(aq) [  $Al^{2*}(aq)$  ] Al(s) $\mathcal{E}^{\bullet} = -1.66 \lor -(-2.37 \lor) = 0.71 \lor$ 

d) Pb & Zn

Cathode: Pb, reaction: Pb<sup>2\*</sup>(aq) + 2e  $\rightarrow$  Pb(s) Anode: Zn, reaction: Zn(s)  $\rightarrow$  Zn<sup>2\*</sup>(aq) +2e Cell notation: Zn(s) | Zn<sup>2\*</sup>(aq) || Pb<sup>2\*</sup>(aq) || Pb(s)  $\mathcal{E}^{\bullet} = -0.13 \text{V} - (-0.76 \text{V}) = 0.63 \text{V}$ 

e) Mg & Cu

Cathode: Cu, reaction:  $Cu^{2*}(aq) + 2e \rightarrow Cu(s)$ Anode: Mg, reaction: Mg(s)  $\rightarrow$  Mg<sup>2\*</sup>(aq) +2e Cell notation: Mg(s) |Mg<sup>2\*</sup>(aq) ||  $Cu^{2*}(aq)$  | Cu(s) $E^0 = 0.34V - (-2.37V) = 2.71V$ 

# **Electrochemistry Exercise Answers**

**RD Boyd** 

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