



## Unit: Stoichiometry **Answer key** "Mole-Mole Calculations" - WS #1

**Directions:** For each of the following problems, show your balanced chemical equation and all of your calculations.

1. How many moles of potassium oxide ( $K_2O$ ) will be formed when 1.52 moles of potassium reacts with oxygen according to the following reaction:



$$1.52 \text{ mol } K \times \frac{2 \text{ mol } K_2O}{4 \text{ mol } K} = 0.760 \text{ mol } K_2O$$

How many moles of copper are needed to react with sulfur to produce 0.25 moles of copper(I) sulfide according to the following reaction:



$$0.25 \text{ mol } Cu_2S \times \frac{2 \text{ mol } Cu}{1 \text{ mol } Cu_2S} = 0.50 \text{ mol } Cu$$

2. How many moles of sodium are needed to react with sulfuric acid to produce 3.75 moles of sodium sulfate according to the following (unbalanced) reaction:



$$3.75 \text{ mol } Na_2SO_4 \times \frac{2 \text{ mol } Na}{1 \text{ mol } Na_2SO_4} = 7.50 \text{ mol } Na$$

3. How many moles of calcium oxide will be produced when 1.6 moles of Iron(III)oxide reacts with calcium phosphate according to the following (unbalanced) reaction:



$$1.6 \text{ mol } Fe_2O_3 \times \frac{3 \text{ mol } CaO}{1 \text{ mol } Fe_2O_3} = 4.8 \text{ mol } CaO$$

4. How many moles of oxygen are produced when 5.12 moles of magnesium chlorate decomposes into magnesium chloride and oxygen?

**Word equation:** **magnesium chlorate  $\rightarrow$  magnesium chloride + oxygen**

**Formula equation:**  **$Mg(ClO_3)_2 \rightarrow MgCl_2 + 3O_2$**   
(must be balanced!)

$$5.12 \text{ mol } Mg(ClO_3)_2 \times \frac{3 \text{ mol } O_2}{1 \text{ mol } Mg(ClO_3)_2} = 15.4 \text{ mol } O_2$$

5. How many moles of Chromium(III) nitrate are produced when Chromium reacts with 0.85 moles of Lead(IV) nitrate to produce chromium(III) nitrate and lead?

**Word equation:** **chromium + lead (IV) nitrate  $\rightarrow$  chromium (III) nitrate + lead**

**Formula equation:**  **$4Cr + 3Pb(NO_3)_4 \rightarrow 4Cr(NO_3)_3 + 3Pb$**   
(must be balanced!)

$$0.85 \text{ mol } Pb(NO_3)_4 \times \frac{4 \text{ mol } Cr(NO_3)_3}{3 \text{ mol } Pb(NO_3)_4} = 1.1 \text{ mol } Cr(NO_3)_3$$

# Moles And Stoichiometry Practice Problems Answer Key

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**Basics for Chemistry** David A. Ucko, 2013-09-24 Basics of Chemistry provides the tools needed in the study of General Chemistry such as problem solving skills calculation methods and the language and basic concepts of chemistry The book is designed to meet the specific needs of underprepared students Concepts are presented only as they are needed and developed from the simple to the complex The text is divided into 18 chapters each covering some particular aspect of chemistry such as matter energy and measurement the properties of atoms description of chemical bonding study of chemical change and nuclear and organic chemistry Undergraduate students will find the book as a very valuable academic material *Class 9: Daily Practice Problems for NTSE, NEET & JEE Foundation (All in One)* Career Point Kota, 2021-12-14 Career Point Kota is one of the first institutes of the country to start DPP concepts for its classrooms students considering the daily practice requirement of the students Keeping in mind the daily practice needs of the students across the nation at large we have come up with DPP Books integrating Daily Practice Problems Sheets The primary focus of this series is to give gradual and daily practice to students through selected questions So that they learn and understand the subject while the course progresses it help students remain engaged and regular in studies Practice Problems Sheets having specific questions on various topics of the individual chapter ensuring the complete Practice of the chapter It is our strong belief that if students work hard on each of the DPP Sheets he she can improve his her learning and master a subject At Career Point we also follow this book in our Classroom Courses We have tried our best to keep errors out of this book Though we shall be grateful to readers who point out any errors and or make constructive suggestions We wish to utilize the opportunity to place on record our special thanks to all members of the Content Development team for their efforts to create this wonderful book Features of this book Cover all subjects concepts 1700 Topic wise chapter wise questions Prepared by Career Point Kota experts **Chemistry** John McMurry, 1998 Chemistry Neil D. Jespersen, Alison Hyslop, 2021-11-02 Chemistry The Molecular Nature of Matter 8th Edition continues to focus on the intimate relationship that exists between structure at the atomic molecular level and the observable macroscopic properties of matter Key revisions in this edition focus on three areas The deliberate inclusion of more updated real world examples that relate common real world student experiences to the science of chemistry Simultaneously examples and questions have been updated to align them with career concepts relevant to the environmental engineering biological pharmaceutical and medical sciences Providing students with transferable skills with a focus on integrating metacognition and three dimensional learning into the text When students know what they know they are better able to learn and incorporate the material Providing a total solution through New WileyPLUS by fully integrating the enhanced etext with online assessment answer specific responses and additional practice resources The 8th edition continues to emphasize the importance of applying concepts to problem solving to achieve high level learning and increase retention of chemistry knowledge Problems are arranged in an intuitive confidence building order **The Practice**

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**Holt Chemistry** Ralph Thomas Myers, 2004

**Foundations of College Chemistry** Morris Hein, Susan Arena, Cary Willard, 2023 Foundations of College Chemistry 16th edition presents chemistry as a modern vital subject and is designed to make introductory chemistry accessible to all beginning students It is intended for students who have never taken a chemistry course or those who had a significant interruption in their studies but plan to continue with the general chemistry sequence The central focus is to make chemistry interesting and understandable and teach students the problem solving skills they will need This International Adaptation offers new and updated content with improved presentation of all course material It builds on the strengths of previous editions including clear explanations and step by step problem solving The material emphasizes real world applications of chemistry as the authors develop the principles that form the foundation for the further study of chemistry There is new and expanded coverage of polarizing power and polarizability Fajans rules collision number and mean free path abnormal molecular masses and van't Hoff factor and applications of radioactivity

**Prentice Hall Chemistry**, 2000

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**Chemistry** James E. Brady, John R. Holum, 1993-01-21 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping students develop skills in analytical thinking and problem solving Students are taught in a variety of ways to think of skills as tools that can be used to solve complex problems Several aids are included to help focus and inspire student interest frequent reference to common chemicals in commercial products numerous photographs of reactions in chapter practice exercises following worked examples

*General Chemistry Workbook* Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning

objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material      **Stoichiometry Unit Project** Luann Marie Decker,1998      *Organic Chemistry, Or, The Happy Carbon* I. David Reingold,2007 This Is A Course In Organic Chemistry Yikes Isn T That The Killer Course That Sophomores Around The World Dread Why Are They Teaching It To Us Students Taking Our First Chemistry Course How Will We Survive      **Merrill Chemistry** Robert C. Smoot,Smoot,Richard G. Smith,Jack Price,1998      **Chemistry: The Central Science** Theodore L. Brown,H. Eugene LeMay Jr.,Bruce E. Bursten,Catherine Murphy,Patrick Woodward,Steven Langford,Dalius Sagatys,Adrian George,2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation      **Basic Concepts of Chemistry** Leo J. Malone,Theodore O. Dolter,2011-12-27 The 9th edition of Malone s Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter WileyPLUS sold separately from text

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