

## Model 2 – Proposed Calculations for Mass of $\text{NH}_3$ to Mass of $\text{N}_2$

### Toby's Method

$$\frac{x \text{ grams}}{50.0 \text{ g}} = \frac{1 \text{ mole } \text{N}_2}{2 \text{ moles } \text{NH}_3} \rightarrow x = \underline{15} \text{ g } \text{N}_2$$

### Rachel's Method

$$50.0 \text{ g } \text{NH}_3 \times \frac{1 \text{ mole } \text{NH}_3}{17.0 \text{ g } \text{NH}_3} = \underline{2.94} \text{ moles } \text{NH}_3$$

$$\frac{x \text{ mole } \text{N}_2}{\underline{1.5} \text{ mole } \text{NH}_3} = \frac{1 \text{ mole } \text{N}_2}{2 \text{ moles } \text{NH}_3} \rightarrow x = \underline{0.11} \text{ moles } \text{N}_2$$

$$\underline{0.11} \text{ mole } \text{N}_2 \times \frac{28.0 \text{ g } \text{N}_2}{1 \text{ mole } \text{N}_2} = \underline{2.43} \text{ g } \text{N}_2$$

### Jerry's Method

$$50.0 \text{ g } \text{NH}_3 \times \frac{1 \text{ mole } \text{NH}_3}{17.0 \text{ g } \text{NH}_3} \times \frac{1 \text{ mole } \text{N}_2}{2 \text{ moles } \text{NH}_3} \times \frac{28.0 \text{ g } \text{N}_2}{1 \text{ mole } \text{N}_2} = \underline{2.43} \text{ g } \text{N}_2$$

- Model 2 shows three proposed calculations to solve the problem in Question 10. Complete the calculations in Model 2 by filling in the underlined values.
- Which method does not use the mole ratio in an appropriate manner? Explain.  
Toby's method does not correctly use the mole ratio that is required.
- Two of the methods in Model 2 give the same answer. Show that they are mathematically equivalent methods.

$$\begin{aligned} &\text{Jerry's method} \\ &\left\{ \begin{aligned} &50 \times \frac{1}{17} = \frac{1}{2} \times \frac{28}{1} = 2.43 \text{ g } \text{N}_2 \\ &\quad \quad \quad \downarrow \\ &\quad \quad \quad \frac{1}{2} \times \frac{1}{17} \\ &\quad \quad \quad \downarrow \\ &0.11 \times \frac{28}{1} = 2.43 \text{ g } \text{N}_2 \end{aligned} \right. \end{aligned}$$

- Use either Rachel or Jerry's method from Model 2 to calculate the mass of hydrogen needed to make 50.0 g of ammonia.  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$

$$50.0 \text{ g} \left( \frac{1 \text{ mole}}{17.0 \text{ g}} \right) \left( \frac{3 \text{ moles}}{2 \text{ moles}} \right) \left( \frac{1 \text{ g}}{2.016 \text{ g}} \right) = 2.43 \text{ g } \text{H}_2 \text{ required}$$

# Mole Ratio Packet Model 2

**Rachel Sandford**



## Mole Ratio Packet Model 2:

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particularly in the brain testes and sperm      *NASA Technical Note* ,1967      **Chemistry** James N. Spencer,George M. Bodner,Lyman H. Rickard,2010-12-28 Chemistry Structure and Dynamics 5th Edition emphasises deep understanding rather than comprehensive coverage along with a focus on the development of inquiry and reasoning skills While most mainstream General Chemistry texts offer a breadth of content coverage the Spencer author team in contrast focuses on depth and student preparation for future studies The fifth edition is revised in keeping with our commitment to the chemical education community and specifically the POGIL Process Oriented Guided Inquiry Learning Project This text reflects two core principles first that the concepts that are covered are fundamental building blocks for understanding chemistry and second that the concepts should be perceived by the students as being directly applicable to their interests and careers The authors further provide this core coverage using 1 of 3 models data driven chemical theories and student understanding which allows for a more concrete foundation on which students build conceptual understanding      Official Gazette of the United States Patent and Trademark Office ,2000      **Air Pollution Abstracts** ,1976      **Air Pollution Abstracts** United States. Environmental Protection Agency,1976      **Proceedings** ,1999      Adsorption Science And Technology, 2nd Pacific Basin Conference Duong D Do,2000-04-06 This book presents the latest research on adsorption science and technology It covers various aspects of materials solid characterization equilibria kinetics determination and new processes      **Official Gazette of the United States Patent Office** United States. Patent Office,1968      Scientific Basis for Nuclear Waste Management ,1986      **Power Systems and Power Plant Control 1989** U. Ahn,2014-06-05 The control of power systems and power plants is a subject of growing interest which continues to sustain a high level of research development and application in many diverse yet complementary areas such as maintaining a high quality but economical service and coping with environmental constraints The papers included within this volume provide the most up to date developments in this field of research      **The Chemical Components of Tobacco and Tobacco Smoke** Alan Rodgman,Thomas A. Perfetti,2016-04-19 Authored by two longtime researchers in tobacco science The Chemical Components of Tobacco and Tobacco Smoke Second Edition chronicles the progress made from late 2008 through 2011 by scientists in the field of tobacco science The book examines the isolation and characterization of each component It explores developments in pertinent analytical      NASA Tech Briefs ,1989      **Scientific Computing in Chemical Engineering II** Frerich Keil,Wolfgang Mackens,Heinrich Voß,Joachim Werther,2012-12-06 The application of modern methods in numerical mathematics on problems in chemical engineering is essential for designing analyzing and running chemical processes and even entire plants Scientific Computing in Chemical Engineering II gives the state of the art from the point of view of numerical mathematicians as well as that of engineers The present volume as part of a two volume edition covers topics such as the simulation of reactive flows reaction engineering reaction diffusion problems and molecular properties The volume is aimed at scientists practitioners and graduate students in chemical engineering industrial engineering and numerical mathematics      *10 in One Study Package*

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## **Table of Contents Mole Ratio Packet Model 2**

1. Understanding the eBook Mole Ratio Packet Model 2
  - The Rise of Digital Reading Mole Ratio Packet Model 2
  - Advantages of eBooks Over Traditional Books
2. Identifying Mole Ratio Packet Model 2
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mole Ratio Packet Model 2
  - User-Friendly Interface
4. Exploring eBook Recommendations from Mole Ratio Packet Model 2
  - Personalized Recommendations
  - Mole Ratio Packet Model 2 User Reviews and Ratings
  - Mole Ratio Packet Model 2 and Bestseller Lists
5. Accessing Mole Ratio Packet Model 2 Free and Paid eBooks
  - Mole Ratio Packet Model 2 Public Domain eBooks
  - Mole Ratio Packet Model 2 eBook Subscription Services
  - Mole Ratio Packet Model 2 Budget-Friendly Options
6. Navigating Mole Ratio Packet Model 2 eBook Formats

- ePub, PDF, MOBI, and More
- Mole Ratio Packet Model 2 Compatibility with Devices
- Mole Ratio Packet Model 2 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Mole Ratio Packet Model 2
  - Highlighting and Note-Taking Mole Ratio Packet Model 2
  - Interactive Elements Mole Ratio Packet Model 2
- 8. Staying Engaged with Mole Ratio Packet Model 2
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Mole Ratio Packet Model 2
- 9. Balancing eBooks and Physical Books Mole Ratio Packet Model 2
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Mole Ratio Packet Model 2
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Mole Ratio Packet Model 2
  - Setting Reading Goals Mole Ratio Packet Model 2
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mole Ratio Packet Model 2
  - Fact-Checking eBook Content of Mole Ratio Packet Model 2
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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