

THE MOLE AND AVOGADRO'S NUMBER

Name KEY

One mole of a substance contains Avogadro's Number (6.02×10^{23}) of molecules.

How many molecules are in the quantities below?

1. 2.0 moles	$\frac{2.0 \text{ mole}}{1 \text{ mole}} \times 6.02 \times 10^{23} \text{ molecules}$	$= 1.20 \times 10^{24} \text{ molecules}$
2. 1.5 moles	$\frac{1.5 \text{ mole}}{1 \text{ mole}} \times 6.02 \times 10^{23} \text{ molecules}$	$= 9.03 \times 10^{23} \text{ molecules}$
3. 0.75 mole	$\frac{0.75 \text{ mole}}{1 \text{ mole}} \times 6.02 \times 10^{23} \text{ molecules}$	$= 4.52 \times 10^{23} \text{ molecules}$
4. 15 moles	$\frac{15 \text{ mole}}{1 \text{ mole}} \times 6.02 \times 10^{23} \text{ molecules}$	$= 9.03 \times 10^{24} \text{ molecules}$
5. 0.35 mole	$\frac{0.35 \text{ mole}}{1 \text{ mole}} \times 6.02 \times 10^{23} \text{ molecules}$	$= 2.11 \times 10^{23} \text{ molecules}$

How many moles are in the number of molecules below?

1. 6.02×10^{23}	$\frac{6.02 \times 10^{23} \text{ molecules}}{6.02 \times 10^{23} \text{ molecules}} = 1 \text{ mole}$	(By definition!)
2. 1.204×10^{24}	$\frac{1.204 \times 10^{24} \text{ molecules}}{6.02 \times 10^{23} \text{ molecules}}$	$= 2.00 \text{ moles}$
3. 1.5×10^{20}	$\frac{1.5 \times 10^{20} \text{ molecules}}{6.02 \times 10^{23} \text{ molecules}}$	$= 2.49 \times 10^{-4} \text{ mole}$
4. 3.4×10^{26}	$\frac{3.4 \times 10^{26} \text{ molecules}}{6.02 \times 10^{23} \text{ molecules}}$	$= 565 \text{ moles}$
5. 7.5×10^{19}	$\frac{7.5 \times 10^{19} \text{ molecules}}{6.02 \times 10^{23} \text{ molecules}}$	$= 1.24 \times 10^{-4} \text{ moles}$

Mole And Avogadros Number Answer Key

Ray Arthur Speckman



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Publishing Fundamentals Sarah S. O'Keefe, 2008-06

Modern Thermodynamics for Chemists and Biochemists

Dennis Sherwood, Paul Dalby, 2018-05-11 Thermodynamics is fundamental to university and college curricula in chemistry physics engineering and many life sciences around the world It is also notoriously difficult for students to understand learn and apply What makes this book different and special is the clarity of the text The writing style is fluid natural and lucid and everything is explained in a logical and transparent manner Thermodynamics is a deep and important branch of science and this book does not make it easy But it does make it intelligible This book introduces a new Fourth Law of Thermodynamics based on the notion of Gibbs free energy which underpins almost every application of thermodynamics and which the authors claim is worthy of recognition as a law The last four chapters bring thermodynamics into the twenty first century dealing with bioenergetics how living systems capture and use free energy macromolecule assembly how proteins fold and macromolecular aggregation how for example virus capsids assemble This is of great current relevance to students of biochemistry biochemical engineering and pharmacy and is covered in very few other texts on thermodynamics The book also contains many novel and effective examples such as the explanation of why friction is irreversible the proof of the depression of the freezing point and the explanation of the biochemical standard state *Chemical Engineering Progress*, 1950

Factors Influencing Microbial Formation of Diacetyl Ray Arthur Speckman, 1980

Extended Abstracts Electrochemical

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