

# Chemistry: Molarity



Directions: Using these formulas, solve the following problems:

molarity = mol solute / L solution

1 amu = 1 gram / mol

$M_1V_1 = M_2V_2$

Problem 1: Calculate the molarity of a 2.21 L solution made with 31.7 g of calcium chloride. Answer: \_\_\_\_\_

Problem 2: How many grams of NaCl would be required to produce a 6.1 M solution with a volume of 0.64 L? Answer: \_\_\_\_\_

Problem 3: How many grams of NaOH would be required to produce a 5.36 M solution with a volume of 1.37 L? Answer: \_\_\_\_\_

Problem 4: How many grams of KCl would you need to weigh out to prepare 0.416 L of a 3.37 M KCl solution? Answer: \_\_\_\_\_

Problem 5: Calculate the molarity of a 6.68 mL solution of 6.89 g NaCl. Answer: \_\_\_\_\_

Problem 6: Calculate the volume in mL if you have 10.9 g of  $\text{Na}_2\text{CO}_3$  from a 0.112 M solution. Answer: \_\_\_\_\_

Problem 7: Calculate the volume in mL if you have 28.1 g LiOH from a 7.11 M solution. Answer: \_\_\_\_\_

Problem 8: Calculate the final concentration if water is added to 0.789 L of a 3.68 M HCl solution to give a volume of 0.387 L. Answer: \_\_\_\_\_

Problem 9: Calculate the final concentration if a 8.63 mL of a 1.12 M HCl solution is diluted with water to 2.54 mL. Answer: \_\_\_\_\_

Problem 10: Calculate the final concentration if water is added to 0.235 L of a 6.11 M KBr solution to give a volume of 5.79 L. Answer: \_\_\_\_\_

# Molarity Worksheet Answers Page 273

**Sasha Hu**



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